

A student investigates the evaporation of water. He pours 100cm^3 of water into measuring cylinder B. Measuring cylinder A is kept at 40°C and B is kept at 80°C in the same part of the laboratory. Fig. 2.1 shows the two measuring cylinders after 3 days.



Fig. 2.1

- (a) (i) Using ideas about molecules, explain why more water evaporates when the water is kept at 80°C .

.....

 [2]

- (ii) Apart from an increase in temperature, state one change that causes water to evaporate faster.

.....
 [1]

- (b) The specific latent heat of vaporisation of water is 2300J/g and the density of water is 1.0g/cm^3 . During the three days, the water level in B drops from the 100cm^3 mark to the level shown in Fig. 2.1. Calculate the energy used to evaporate water from B during the three days.

energy = [2]

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L Reisser



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Table of Contents Igcse Physics May 2013 Question Paper 12

1. Understanding the eBook Igcse Physics May 2013 Question Paper 12
 - The Rise of Digital Reading Igcse Physics May 2013 Question Paper 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Physics May 2013 Question Paper 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Physics May 2013 Question Paper 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Physics May 2013 Question Paper 12
 - Personalized Recommendations
 - Igcse Physics May 2013 Question Paper 12 User Reviews and Ratings
 - Igcse Physics May 2013 Question Paper 12 and Bestseller Lists
5. Accessing Igcse Physics May 2013 Question Paper 12 Free and Paid eBooks

- Igcse Physics May 2013 Question Paper 12 Public Domain eBooks
 - Igcse Physics May 2013 Question Paper 12 eBook Subscription Services
 - Igcse Physics May 2013 Question Paper 12 Budget-Friendly Options
6. Navigating Igcse Physics May 2013 Question Paper 12 eBook Formats
 - ePub, PDF, MOBI, and More
 - Igcse Physics May 2013 Question Paper 12 Compatibility with Devices
 - Igcse Physics May 2013 Question Paper 12 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Igcse Physics May 2013 Question Paper 12
 - Highlighting and Note-Taking Igcse Physics May 2013 Question Paper 12
 - Interactive Elements Igcse Physics May 2013 Question Paper 12
 8. Staying Engaged with Igcse Physics May 2013 Question Paper 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Physics May 2013 Question Paper 12
 9. Balancing eBooks and Physical Books Igcse Physics May 2013 Question Paper 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Physics May 2013 Question Paper 12
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Igcse Physics May 2013 Question Paper 12
 - Setting Reading Goals Igcse Physics May 2013 Question Paper 12
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Igcse Physics May 2013 Question Paper 12
 - Fact-Checking eBook Content of Igcse Physics May 2013 Question Paper 12
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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