

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages created through Eureka, Slides, and Animations and answer the following questions in detail and understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

Concept Name	Bohr Model	Quantum Model
Electrons	positively charged particles in the atom	positively charged particles in the atom
Orbitals	negatively charged and orbit the nucleus	negatively charged clouds
Location of the electron in the atom without	is in two fixed paths	electron cloud pattern
Energy of the electron in the atom without	electron has fixed energy levels	energy has probability
Energy of the electron in the atom without	$E = -13.6/n^2$	$E = -13.6/n^2$

Question 2: Complete the following table which shows the comparison of two different theories by putting predictions made by each theory side-by-side. The table can be completed by using the Bohr model or Quantum model. Complete the table by using the Bohr model or Quantum model. Complete the table by using the Bohr model or Quantum model. Complete the table by using the Bohr model or Quantum model.

Hydrogen Atom Student Guide Naap Answers

**Rudolph M. Langer, California Institute
of Technology. Division of Physics,
Mathematics and Astronomy**



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