

NUMERICALS (CHEM-I)



Dr. ZAIN UL ABDEEN

CHAPTER # 2

GRAHAM'S LAW OF DIFFUSION

QUESTION # 24, pg. 72, STBB

Helium gas takes 5 seconds to effuse from a hole of 10 dm³ container, how long it would take for 10 dm³ of oxygen gas to effuse from the same container at the same temperature and pressure?

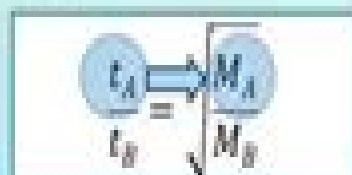
DATA:

$$t_{He} = 5 \text{ s}$$

$$M_{He} = 4 \text{ amu}$$

$$M_{O_2} = 32 \text{ amu}$$

$$t_{O_2} = ?$$



Direct Relation

Finding t_{O_2} :

Hint: (The substance which is required should taken as numerator)

Using Graham's Law

$$\frac{t_{O_2}}{t_{He}} = \sqrt{\frac{M_{O_2}}{M_{He}}}$$

$$\frac{t_{O_2}}{5} = \sqrt{\frac{32}{4}}$$

$$\frac{t_{O_2}}{t_{He}} = \sqrt{8}$$

$$\frac{t_{O_2}}{5} = 2.828$$

$$t_{O_2} = 2.828 \times 5$$

$$t_{O_2} = 14.14 \text{ s}$$

Ends...

Grahams Law Of Diffusion Answer Key

Mary Millhollon, Richard H. Langley



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Table of Contents Grahams Law Of Diffusion Answer Key

1. Understanding the eBook Grahams Law Of Diffusion Answer Key
 - The Rise of Digital Reading Grahams Law Of Diffusion Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Grahams Law Of Diffusion Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grahams Law Of Diffusion Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grahams Law Of Diffusion Answer Key
 - Personalized Recommendations
 - Grahams Law Of Diffusion Answer Key User Reviews and Ratings
 - Grahams Law Of Diffusion Answer Key and Bestseller Lists
5. Accessing Grahams Law Of Diffusion Answer Key Free and Paid eBooks
 - Grahams Law Of Diffusion Answer Key Public Domain eBooks
 - Grahams Law Of Diffusion Answer Key eBook Subscription Services

- Grahams Law Of Diffusion Answer Key Budget-Friendly Options
- 6. Navigating Grahams Law Of Diffusion Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Grahams Law Of Diffusion Answer Key Compatibility with Devices
 - Grahams Law Of Diffusion Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grahams Law Of Diffusion Answer Key
 - Highlighting and Note-Taking Grahams Law Of Diffusion Answer Key
 - Interactive Elements Grahams Law Of Diffusion Answer Key
- 8. Staying Engaged with Grahams Law Of Diffusion Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grahams Law Of Diffusion Answer Key
- 9. Balancing eBooks and Physical Books Grahams Law Of Diffusion Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grahams Law Of Diffusion Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grahams Law Of Diffusion Answer Key
 - Setting Reading Goals Grahams Law Of Diffusion Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grahams Law Of Diffusion Answer Key
 - Fact-Checking eBook Content of Grahams Law Of Diffusion Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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