

Qualitative Tests for Hydrocarbons, Alkanes, Alkenes, Alkynes

Hydrocarbons can be grouped into two structural classes and sample chemical tests can be used to distinguish between these classes. In this experiment you will perform four tests: combustion, bromine water, potassium permanganate, and solubility. Each using known hydrocarbons and their derivatives that were used as standards for identifying unknowns observed after.

Five characteristics of hydrocarbons are as follows:

- Alkanes**—There are saturated hydrocarbons, usually single bonds.
- Alkenes**—There are hydrocarbons that contain at least one carbon-carbon double bond.
- Alkynes**—There are hydrocarbons that contain only triple bonds.
- Aromatics**—There are hydrocarbons that contain an aromatic ring. An aromatic ring is a structure composed of alternating double and single bonds that is generally stable due to the delocalization of electrons around the ring. Although the generally accepted structure is two double bonds, other double bonds are delocalized in the representation. They are small structures that represent double bonds. There are also the fact that they are not really double bonds but a mixture of double and single bonds that is a real molecule itself. There are also some of the following properties given by hydrocarbons:



Put the ball-and-stick model on a white card, the hydrocarbon model as Figure 11.2.



Test the hydrocarbons and hydrocarbon ions

The hydrocarbons of alkanes and alkenes are very easily burned without color and other changes of color. The hydrocarbons of the alkynes are very easily burned with color changes and other changes of color. The hydrocarbons of alkenes are very easily burned with color changes and other changes of color. The hydrocarbons of alkenes are very easily burned with color changes and other changes of color. The hydrocarbons of alkenes are very easily burned with color changes and other changes of color.

Procedure:

Sample 1: alkane, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. Add 1 mL of 10% sodium hydroxide. The solution is colorless and clear. Sample 2: alkene, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 3: alkyne, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 4: aromatic, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 5: alkane, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 6: alkene, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 7: alkyne, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear. Sample 8: aromatic, dry test tube, add about 1 mL of 10% potassium permanganate solution and add some water. The solution is colorless and clear.



Hydrocarbon Lab Manual

JE Gale



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