

Chapter 7 Answers

Practice 7-1

1. 8 2. 8 3. 60 4. 240 5. 2,635 6. 1.92
7. 0.8125 8. 9 9. 1500 10. 8.75 11. 3.44
12. 7 13. 110 14. 30 15. 16 16. 119
17. 12 18. 48 19. 40 20. 56

Practice 7-2

1. $\sqrt{51}$ 2. $4\sqrt{7}$ 3. $2\sqrt{65}$ 4. $\sqrt{7}$
5. $2\sqrt{21}$ 6. $18\sqrt{2}$ 7. 210 in.^2 8. $24\sqrt{10} \text{ cm}^2$
9. $35\sqrt{114} \text{ ft}^2$ 10. 46 in. 11. 78 ft 12. 279 cm
13. 19 m 14. acute 15. right 16. obtuse
17. right 18. obtuse 19. acute

Practice 7-3

1. $x = 2; y = \sqrt{3}$ 2. $a = 4.5; b = 4.5\sqrt{3}$
3. $c = \frac{80\sqrt{3}}{3}; d = \frac{20\sqrt{3}}{3}$ 4. $8\sqrt{2}$ 5. $14\sqrt{2}$
6. $\frac{28\sqrt{3}}{3}$ 7. 2 8. $x = 15; y = 15\sqrt{3}$ 9. $3\sqrt{2}$
10. 42 cm 11. 5.9 in. 12. 10.4 ft, 12 ft 13. 170 in.^2
14. $w = \frac{10\sqrt{3}}{3}; x = 5; y = 5\sqrt{2}; z = \frac{2\sqrt{2}}{3}$ 15. $a = 4;$
 $b = 3$ 16. $p = 4\sqrt{3}; q = 4\sqrt{3}; r = 8; s = 4\sqrt{6}$

Practice 7-4

1. 48 cm^2 2. 784 in.^2 3. 11.4 ft^2 4. 90 m^2
5. 2400 in.^2 6. 374 ft^2 7. 160 cm^2 8. 176.25 in.^2
9. 54 ft^2 10. 42.5 square units 11. $96\sqrt{3}$ square units
12. $36\sqrt{3}$ square units 13. 45 cm^2 14. 226.2 in.^2
15. $49,500 \text{ m}^2$

Practice 7-5

1. $x = 7; y = 3.5\sqrt{3}$ 2. $a = 60; c = \frac{8\sqrt{3}}{3}; d = \frac{8\sqrt{3}}{3}$
3. $p = 4\sqrt{3}; q = 8; x = 30$ 4. $m\angle 1 = 45; m\angle 2 = 45;$
 $m\angle 3 = 90; m\angle 4 = 90$ 5. $m\angle 5 = 60; m\angle 6 = 30;$
 $m\angle 7 = 120$ 6. $m\angle 8 = 60; m\angle 9 = 30; m\angle 10 = 60$
7. $12\sqrt{3}$ 8. $36.75\sqrt{3}$ 9. $75\sqrt{3}$ 10. 120 in.^2
11. 137 in.^2 12. 97 in.^2

Practice 7-6

1. 32π 2. 16π 3. 7.8π 4. Samples: $\overline{DFA}, \overline{ABF}$
5. Samples: $\overline{FE}, \overline{FD}$ 6. Samples: $\overline{FEB}, \overline{EDA}$
7. Samples: $\overline{FE}, \overline{ED}$ 8. $\angle ACB$ 9. $\angle FCD$
10. 32 11. 43 12. 54 13. 39 14. 71
15. 90 16. 50 17. 220 18. 140 19. 220
20. 130 21. $6\pi \text{ in.}$ 22. $16\pi \text{ cm}$ 23. $\frac{3}{4}\pi \text{ m}$

Practice 7-7

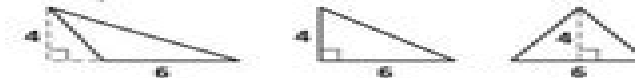
1. 49π 2. 21.2 3. $\frac{92}{5}\pi$ 4. $\frac{92}{5}\pi - 21.2$ 5. $\frac{1}{3}\pi$
or $\frac{\pi}{3}$ 6. $\frac{1}{3}$ 7. $\frac{1}{10}\pi$ or $\frac{\pi}{10}$ 8. $\frac{\pi}{10} - \frac{1}{3}$ 9. 4π
10. $\frac{12}{5}\pi$ 11. $\frac{2}{3}\pi$ 12. $\frac{3}{2}\pi$ 13. $\frac{12}{5}\pi$ 14. 6π
15. $\frac{92}{5}\pi$ 16. $\frac{92}{5}\pi$ 17. $\frac{\pi}{5}$ 18. 3 19. 2
20. 10

Practice 7-8

1. 8.7% 2. 10.9% 3. 15.3% 4. $\frac{3}{10}$ 5. $\frac{7}{10}$
6. $\frac{3}{5}$ 7. $\frac{1}{10}$ 8. 1 9. $\frac{1}{3}$ 10. 22.2% 11. 4.0%
12. 37.5% 13. 20% 14. 50% 15. 40%
16. $33\frac{1}{3}\%$

Reteaching 7-1

1. Samples:

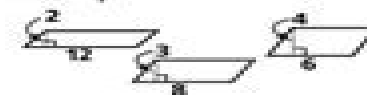


2. Samples:



3. 20 cm^2 4. 55 in.^2 5. 6 m 6. 8 ft 7. 4 in.

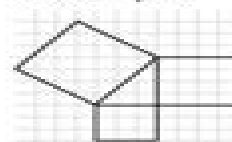
8. Samples:



9. 5 in. 10. 23 cm 11. 9 ft

Reteaching 7-2

1–2. Sample:



3. Sample: 9, 16, and 25 4. They are equal. 5. The sum of the squares of the two legs equals the square of the hypotenuse. 6. $4\sqrt{5}$ 7. $7\sqrt{2}$ 8. 8 9. 6

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G Psacharopoulos



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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of

substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalization of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7.0002g of copper ii sulfate and 2.0101g of iron metal which produced 2.3037g of copper product with a percent yield of 100.72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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