Cinizizz	NAME:	
Bronsted-lowry acids and bases 10 Questions	CLASS:	
	DATE:	

1.40	Question is					
1.	What might happen if you mixed a strong acid with an equally strong base?					
.A.	The acid would destroy	the base.	1	You would see an explosive chemical reaction.		
C	The base would destro	y the acid.		You'd wind up with a pH-neutral substance.		
2.	POL" + HNO. → NO." + HPOL" What is the conjugate acid in the following equation?					
$_{j}\Delta_{i_{j}}$	HPO4 <sup>2</sup>	В	1	PO4 <sup>3-</sup>		
C	HNO <sub>3</sub>	D	1	NO <sub>3</sub> °		
3.	$BCO_3^- + BCI \rightarrow B_2CO_3 + CI^-$ What is the conjugate base in the following reaction?					
,А,	CI_	В	1	HCI		
[0]	H <sub>2</sub> CO <sub>3</sub>	D	1	HCO <sub>3</sub>		
4	$HSO_4$ (aq)+ $H_2O(l)$ — $H_3O^+$ (aq)+ $SO_4^{2-}$ (aq) Identify the Acid in the above reaction.					
A.	SO <sub>4</sub> <sup>2</sup> ·	В	]	H₃O*		
C	HSO <sub>4</sub>	D		H <sub>2</sub> O		
5.	$CO_3^{2}$ ·(aq)+ $H_2O(I) \rightarrow HCO_3$ ·(aq)+ $OH$ ·(aq) Using the above reaction, which compound is the conjugate base?					
,A,	HCO <sub>3</sub>	В	3	OH-		
[C]	H <sub>2</sub> O	D	1	CO3 <sup>2-</sup>		

B H<sub>2</sub>O + lonk Salt

What are the products to a neutralization reaction?

OH'+ Ionic Salt

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