Acids and Bases

How do acids and bases behave in water?

Why?

Acids and bases play an important role in our lives. Numerous biological processes, industrial applications, and even environmental problems are a function of the acidity or basicity (alkalinity) of aqueous solutions. It is therefore important to understand what makes a substance behave as an acid or a base when dissolved in water. In this activity, we will explore the physical and chemical properties of acids and bases.

Model 1 – Arrhenius Acids and Bases

Common Name for Aqueous Solution	Chemical Formula	Found in	Tastes	Turns Litmus Paper	Conducts Electricity?	Acid or Base
Acetic acid	HC,H,O,(aq)	Vinegar	Sour	Red	Yes	Acid
Benzoic acid	HC,H,O,(aq)	Food preservative	Sour	Red	Yes	Acid
Phosphoric acid	H,PO4(aq)	Soda pop	Sour	Red	Yes	Acid
Hydrochloric acid	HCl(aq)	Stomach acid	Sour	Red	Yes	Acid
Citric acid	H ₄ C ₄ H ₄ O ₅ (aq)	Citrus fruits	Sour	Red	Yes	Acid
Ascorbic acid	$H_2C_aH_aO_a(aq)$	Vitamin C	Sour	Red	Yes	Acid
Magnesium hydroxide	Mg(OH),(aq)	Milk of magnesia	Bitter	Blue	Yes	Base
Aluminum hydroxide	Al(OH),(aq)	Antacids	Bitter	Blue	Yes	Basse
Barium hydroxide	Ba(OH),(aq)	Lubricants	POISON	Blue	Yes	Base
Sodium hydroxide	NaOH(aq)	Drain cleaner	POISON	Blue	Yes	Base

- 1. Refer to Model 1.
 - What is the common chemical name for Vitamin C? ascorbic acid
 - b. Is Vitamin C classified as an acid or a base? acid
- Examine the properties of the Arrhenius acids in Model 1. List three properties that all Arrhenius acids have in common.

sour, red, conduct electricity

- 3. Examine the chemical formulas for the Arrhenius acids in Model 1. What feature do all the Arrhenius acid chemical formulas have in common?
- Examine the properties of the Arrhenius bases in Model 1. List two properties that all Arrhenius bases have in common.

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IM Harris

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