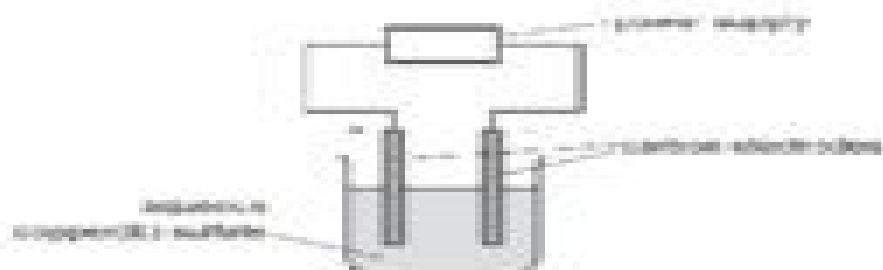


**Question 1 \* 64/66 (41/48/12/00)**

(a) A student electrolyses aqueous copper(II) sulfate using the apparatus shown.



Oxygen gas forms at the positive electrode (anode).

(b) Write an ionic half-equation for the reaction at the negative electrode (cathode). Include state symbols.

\_\_\_\_\_ [1]

(c) Describe what the student observes at the negative electrode.

\_\_\_\_\_ [1]

(d) Give two other observations which the student makes during the electrolysis.

1. \_\_\_\_\_

2. \_\_\_\_\_

[2]

(e) What difference would the student observe at the positive electrode if the aqueous copper(II) sulfate were replaced by concentrated aqueous copper(II) chloride?

\_\_\_\_\_ [1]

**Question 2 \* 64/66 (40/48/12/02/14)**

(a) Electrolysis of concentrated aqueous sodium chloride is an important industrial process.

(i) What is meant by the term anode/cathode?

\_\_\_\_\_

\_\_\_\_\_ [1]

# Igcse Electrolysis Past Papers

**Julia Schneider**



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