

9. A 12.0 L sample of nitrogen has a pressure of 540 kPa at 150 °C. The gas is compressed to 1.20 L and the temperature is lowered to 50 °C. What is the new pressure?

$$\frac{540 \text{ kPa} \cdot 12.0 \cancel{\text{L}}}{423 \text{ K}} = \frac{P_2 \cdot 1.20 \cancel{\text{L}}}{323 \text{ K}}$$

$$P_2 = 4,123 \text{ kPa}$$

10. What mass of helium is needed to inflate a balloon to 4.0 L tank at 305 K and 2.5 atm?

$$n = \frac{PV}{RT} = \frac{(2.5 \text{ atm})(4.0 \cancel{\text{L}})}{(0.0821 \text{ atm} \cdot \text{L} / \text{mol} \cdot \text{K})(305 \text{ K})} = 0.399 \text{ mol He} \left| \frac{4 \text{ g He}}{1 \text{ mol He}} \right| = \boxed{1.6 \text{ g He}}$$

11. Compare the diffusion rates of helium and radon.

12. A sample of ozone is collected over water at 27 °C and at a barometric pressure of 765 torr. What is the partial pressure of the ozone gas?

$$P_{O_2} = 765 \text{ torr} - 26 \text{ torr} = 739 \text{ torr}$$

13. Are volume and temperature inversely or directly related? Give a real life example.

$\frac{V}{T}$ As Temp. inc., vol. inc.
when pouring hot water over balloon, balloon expands

14. A mixture of gases contains N₂, O₂, and CO₂. If the total pressure of the mixture is 10.0 atm and N₂ = 2.0 atm and O₂ = 3.5 atm, what is the partial pressure of CO₂?

$$10.0 \text{ atm} = 2.0 \text{ atm} + 3.5 \text{ atm} + P_{O_2}$$

$$(4.5 \text{ atm} = P_{O_2})$$

15. If a sample contains 8 moles of H₂, 3 moles of O₂, and 2 moles of He in a 7 liter vessel at 19 °C, determine the partial pressure of each gas and the total pressure of the mixture in units of atm.

$$P_{H_2} = \frac{(8 \text{ mol H}_2)(0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K})(292 \text{ K})}{7 \text{ L}} = 27.4 \text{ atm}$$

$$P_{O_2} = \frac{(3 \text{ mol O}_2)(0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K})(292 \text{ K})}{7 \text{ L}} = 10.3 \text{ atm}$$

$$P_{He} = \frac{(2 \text{ mol He})(0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K})(292 \text{ K})}{7 \text{ L}} = 6.84 \text{ atm}$$

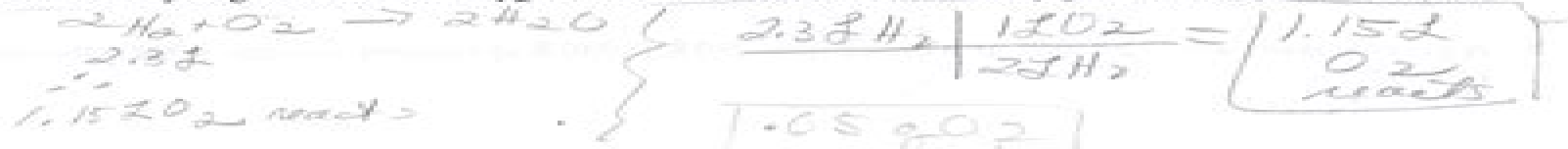
$$P_T = 44.5 \text{ atm}$$

16. If a gas sample with a pressure of 86.2 kPa contains 53% methane gas, what is the partial pressure of methane?

$$86.2 \text{ kPa} = 45.7 \text{ kPa} + 40.5 \text{ other gas}$$

$$P_T = 86.2 \text{ kPa} \quad P_{CH_4} = (0.53)(86.2 \text{ kPa}) = 45.7 \text{ kPa}$$

17. If 2.3 L of hydrogen reacts with oxygen to form water at STP, how many grams of oxygen are reacting?



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A Gutmann



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