

Empirical/Molecular Formula

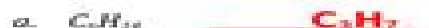
Name: _____

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1. What is an empirical formula? How is it different from a molecular formula?

The empirical formula is the lowest whole number ratio of atoms in a compound. The molecular formula shows the true number of atoms in the molecule and is a whole number multiple of the empirical formula.

2. State the empirical formula for each of the following.



3. A sample of a compound is 50.0% sulfur and 50.0% oxygen by mass. What is the empirical formula?

$$S: \frac{50.0 \text{ g S}}{32.07 \text{ g S}} \times \frac{1 \text{ mol S}}{1} = \frac{1.559 \text{ mol S}}{1.559 \text{ mol}} = 1 \text{ S}$$

EF: SO_2

$$O: \frac{50.0 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{3.125 \text{ mol O}}{1.559 \text{ mol}} = 2 \text{ O}$$

4. A sample of a compound is 52.9% aluminum. The rest is oxygen. What is the empirical formula?

$$Al: \frac{52.9 \text{ g Al}}{26.98 \text{ g Al}} \times \frac{1 \text{ mol Al}}{1} = \frac{1.961 \text{ mol Al}}{1.961 \text{ mol}} = 1 \text{ Al} \times 2 = 2$$

EF: Al_2O_3

$$O: \frac{47.1 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{2.944 \text{ mol O}}{1.961 \text{ mol}} = 1.5 \times 2 = 3 \text{ O}$$

Ratio must be a whole

5. A compound is 48.65% carbon, 8.11 % hydrogen and 43.24 % oxygen. Determine the empirical formula.

4.05 mol carbon, 8.11 mol H, 2.70 mol O



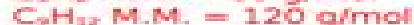
1. Divide by the smallest # to get = $C_{1.5}H_3O_1$

2. Multiply to get a whole number ratio (x2) = $C_3H_6O_2$

6. Vinegar is a dilute solution of acetic acid. The molar mass of acetic acid is 60.0 g/mol and it has an empirical formula of CH_2O . What is the molecular formula of acetic acid?



7. A compound has an empirical formula of C_3H_4 . Which of the following are possible molar masses of the compound? 20 g/mol, 55 g/mol, 80 g/mol, 120 g/mol.



Empirical And Molecular Formula Show Work

Scott C. Dulebohn



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