The empirical formula of a compound represents the lowest whole number male ratio
of component elements.

The molecular formula is the actual ratio of elements found with the naturally-occurring sample. This could also equal the empirical formula if the naturally-occurring formula is also in the lowest whole number mole ratio.

- Molecular = CataNaO + each subscript can be reduced by 2 Empirical = CABNAO
- 4) Compound pair B is the only set of compounds with the same empirical formulas.

 To determine the empirical formula, convert each atom's provided mass into moles and find the lowest whole number mole ratio.

```
\begin{array}{c} 0.8007g~C\times1~mol~C/12.01g~C=0.066669~mol~C\\ 0.9333g~N\times1~mol~N/14.01g~N=0.066617~mol~N\\ 0.2016g~H\times1~mol~N/1.000g~H=0.2~mol~H\\ 0.2133g~O\times1~mol~0/16.00g~O=0.133313~mol~O\\ \end{array}
```

6) Determine the empirical formula first, by assuming that 100g of the sample is available, allowing for 24.27% C ÷ 24.27g C, etc... Use the relative masses of each element to determine moles and find the lowest whole number mole ratio. Then compare the mass of the empirical formula to that of the mass of the molecular formula provided. This comparison will allow you to determine the molecular formula.

Molecular mass/emointed mass = n - n/empirical formula\(\) = molecular

```
24.27% C → 24.27 g C × 1 mol C/12.01g C = 2.021 mol C

6.07% H → 4.07g H × 1 mol Hy1.008g H = 4.037 mol H

71.65% Cl → 71.65g Cl × 1 mol Cl/35.45 g Cl = 2.021 mol Cl

2.021 mol C/2.021 = 1 mol C

4.037 mol Hy2.021 = 2 mol H} CH_Cl = empirical formula

2.021 mol Cl/2.021 = 1 mol C

Empirical mass of CH_Cl → Cl 12.01g +

Hz 201.008gl +

Cl: 35.45 g

To compare empirical & molecular mass:
```

formeday.

98,96 g/49,476 g - 2

MICHAEL - CHACL

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Heather Hattori, Richard H. Langley

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