

Section A

Answer all questions in the spaces provided

0 1

Anhydrous magnesium chloride, MgCl_2 , can absorb water to form the hydrated salt $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$



← creating crystals directly is not really possible.

0 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

0 1 - 2

Some enthalpies of solution are shown in Table 1.

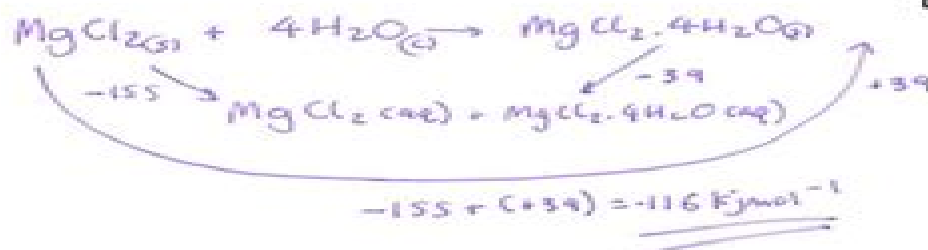
Table 1

Salt	Enthalpy of solution / kJ mol^{-1}
$\text{MgCl}_2(\text{s})$	-155
$\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$	-39

making the salts aqueous!

Calculate the enthalpy change for the absorption of water by $\text{MgCl}_2(\text{s})$ to form $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$.

[2 marks]



Enthalpy change _____ kJ mol^{-1}



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Table of Contents Max Papers Chemistry Alevel 2014

1. Understanding the eBook Max Papers Chemistry Alevel 2014
 - The Rise of Digital Reading Max Papers Chemistry Alevel 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Max Papers Chemistry Alevel 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Max Papers Chemistry Alevel 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Max Papers Chemistry Alevel 2014
 - Personalized Recommendations
 - Max Papers Chemistry Alevel 2014 User Reviews and Ratings
 - Max Papers Chemistry Alevel 2014 and Bestseller Lists
5. Accessing Max Papers Chemistry Alevel 2014 Free and Paid eBooks

- Max Papers Chemistry Alevel 2014 Public Domain eBooks
- Max Papers Chemistry Alevel 2014 eBook Subscription Services
- Max Papers Chemistry Alevel 2014 Budget-Friendly Options
- 6. Navigating Max Papers Chemistry Alevel 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Max Papers Chemistry Alevel 2014 Compatibility with Devices
 - Max Papers Chemistry Alevel 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Max Papers Chemistry Alevel 2014
 - Highlighting and Note-Taking Max Papers Chemistry Alevel 2014
 - Interactive Elements Max Papers Chemistry Alevel 2014
- 8. Staying Engaged with Max Papers Chemistry Alevel 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Max Papers Chemistry Alevel 2014
- 9. Balancing eBooks and Physical Books Max Papers Chemistry Alevel 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Max Papers Chemistry Alevel 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Max Papers Chemistry Alevel 2014
 - Setting Reading Goals Max Papers Chemistry Alevel 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Max Papers Chemistry Alevel 2014
 - Fact-Checking eBook Content of Max Papers Chemistry Alevel 2014
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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