Acid and Base Worksheet - Answers

- Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair.
 - a) HNO₂ + OH¹ → H₂O + NO₃¹ HNO₂ and NO₃¹ make one pair OH¹ and H₂O make the other
 - b) CH₃NH₂ + H₂O → CH₃NH₃* + OH¹
 CH₃NH₂ and CH₃NH₃* make one pair OH¹ and H₂O make the other
 - c) OH¹ + HPO₄² → H₂O + PO₄³ HPO₄² and PO₄³ make one pair OH¹ and H₂O make the other
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
 - NaOH is an Arrhenius base because it creates OH¹ ions when placed in water.
 - NaOH is a Brønsted-Lowry base because it accepts H^{*1} ions from acids.
 - NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.
- When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
 HCl + NH₃ → NH₄Cl
 NH₃ is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.
 HCl is a Lewis acid because it accepts electrons from NH₃ when the H is transferred.
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
 - Carbon dioxide and water
 CO₂ + H₂O → H₂CO₃
 Acid anhydride
 - b. Sodium oxide and water
 Na₂O + H₂O → 2NaOH(aq)
 Basic anhydride
 - c. Dichlorine trioxide and water
 Cl₂O₂ + H₂O → 2HClO₂
 Acidic anhydride
 - d. Beryllium oxide and water
 BeO + H₂O → Be(OH)₂
 Basic anhydride

Introduction To Acids And Bases Answer Key

Pascal Van Der Voort, Karen Leus, Els De Canck

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