



**11** How does the arrangement of particles in a solid differ from the arrangement of particles in a liquid? What happens when a solid melts? What happens when a liquid freezes? What happens when a gas condenses? What happens when a liquid evaporates? What happens when a solid sublimates? What happens when a gas condenses? What happens when a liquid evaporates? What happens when a solid sublimates?

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**9** What is the difference between a solid, a liquid, and a gas? What are the properties of each state of matter?

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## PARTICLE THEORY AND STATES OF MATTER

Particle theory explains the properties of matter in terms of the arrangement of particles. The particles in matter are constantly moving, and the arrangement of particles determines the state of matter. The particles in a solid are closely packed together, while the particles in a liquid are more spread out. The particles in a gas are very spread out and move rapidly.

**States of Matter:** Matter can exist in three primary states: solid, liquid, and gas. Each state has different properties. Solids have a fixed shape and volume. Liquids have a fixed volume but no fixed shape. Gases have no fixed shape or volume.

**Changes of State:** Matter can change from one state to another. These changes are called changes of state. The most common changes of state are melting, freezing, boiling, and condensation.

**Particle Theory and Changes of State:** Changes of state occur because the particles in matter are constantly moving. When a solid melts, the particles gain enough energy to break the bonds that hold them together. When a liquid freezes, the particles lose enough energy to form bonds again.

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# Introducing To Matter Guided Reading Answer Key

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