

Intro to Acids and Bases

Name: _____
 Class: _____
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Define each term: Definitions of acids and bases:

Arrhenius-definitions:

- acid substances that release dissolved cations hydrogen (H^+) (actual H^+ cations bonded to water as the hydronium ion, H_3O^+ , and H^+ (actual H_2O^+) and hypothetical equivalents)
- base substances which release dissolved water hydroxide (OH^-)

Brønsted-Lowry definitions:

- acid substances that donate a proton (H^+) (also accepted)
- base substances which accept a proton (H^+) (also donated)

Table 1: Representative reactions for H^+ (also known as the Lewis acid) and bases according to Arrhenius:

Reaction	Eq.
1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$	$H^+(aq) + H_2O(l)$
2. $HNO_3 + H_2O \rightleftharpoons H_3O^+ + NO_3^-$ (actual H^+ is H_3O^+)	$H^+ + H_2O(l)$
3. $HNO_3(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + NO_3^-(aq)$	$H^+(aq) + H_2O(l)$
4. $H_2SO_4 + H_2O(l) \rightleftharpoons H_3O^+ + HSO_4^-$	$H^+(aq) + H_2O(l)$
5. $HCl + H_2O(l) \rightleftharpoons H_3O^+ + Cl^-$	$H^+ + H_2O(l)$

Critical Thinking Questions

1. List acid characteristics for Table 1 that are Arrhenius acids:

aq, $HCl(aq)$, $HNO_3(aq)$, H_2SO_4

2. List acid characteristics for Table 1 that are Arrhenius bases:

$Cl^-(aq)$

3. List acid characteristics in Table 1 that are Arrhenius acids and bases:

$H_2O(l)$, $H_3O^+(aq)$, $NO_3^-(aq)$, $HSO_4^-(aq)$, $Cl^-(aq)$

4. List acid characteristics in Table 1 that are Arrhenius acids and bases:

$HCl(aq)$, $HNO_3(aq)$

5. Is it possible for an acid to be an Arrhenius base? Explain:

Yes, because $H_2O(l)$, $H_3O^+(aq)$ can be both.

6. List acid characteristics in Table 1 that are Arrhenius acids and bases:

aq, $HCl(aq)$, $HNO_3(aq)$, H_2SO_4 , $H_3O^+(aq)$, $NO_3^-(aq)$, $HSO_4^-(aq)$, $Cl^-(aq)$

7. Is it possible for an acid to be an Arrhenius base? Explain:

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