- 29. For a cell Ag|Ag⁺,Cl[−] |AgCl(s)|Ag, E^s_{Red} for the half cells are given as: E^s_{Ag⁺|Ag} = 0.79 V, E^s_{Cl[−]|AgCl(s)|Ag} = 0.22 V at 25 °C. What is the value of lnK, where K is the equilibrium constant for the reaction AgCl(s) = Ag⁺ + Cl[−]? Given, 1 F = 96485 C.
 - A. -22.2
 - $B_{\rm e} = 18.5$
 - C. -29.3
 - D. -26.8
- Match the following:
 - Ge doped with In 1. n-type semiconductor
 - ii) Si doped with N 2. Schottky defect
 - iii) ZnS 3. p-type semiconductor
 - iv) CsCl 4. Frenkel defect
 - A. i-1, ii-3, iii-2, iv-4
 - B. i-1, ii-3, iii-4, iv-2
 - C, i-3, ii-1, iii-2, iv-4
 - D. i-3, ii-1, iii-4, iv-2

<u>Iiser Aptitude Test 20question Paper</u>

Michael Rajnik

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