Use noble gases to wr	te the abbreviated electron configurations of the follow
i. Li	
iii. O	iv. Mg
v. Cl	vi. V
vii. Se ²⁻	viii. Cr²+
ix. I	x. Ag
2. Identify the element v	ith the following electron configuration
i. 1s ² 2s ² 2p° 3s ² 3p° 4s ² 3c	10 4p4
ii. 1s² 2s² 2p° 3s² 3p° 4s² 3	1 ¹⁰ 4p ⁶ 5s ²
iii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	d ¹⁰ 4p ⁶ 5s ¹ 4d ⁷
iv. 1s2 2s2 2p4 3s2 3p4 4s2 3	d ¹⁰ 4p ⁶ 5s ² 4d ¹⁰ 5p ⁶ 6s ² 4f ¹⁴ 5d ¹⁰
v. 1s2 2s2 2p6 3s2 3p6 4s2 3	d ¹⁰ 4p° 5s² 4d ¹⁰ 5p° 6s² 4f ¹⁴ 5d ¹⁰ 6p° 7s¹
3. Determine if the follo i. 1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ² 4c ² ii. 1s ² 2s ² 2p ⁶ 3s ³ iii. [Rn] 7s ² 5f ⁶ 6d ² iv. [Ar] 5s ² 4d ¹⁰ 5p ³ v. [Xe] 6s ² 4f ¹⁰	ving electron configurations are correct: (10 4p ⁶ 5s ¹
4. Identify the following	elements
i. 1s ² 2s ² 2p ²	
ii. 1s² 2s² 2p°	
iii. [Ar] 4s2 3d30 4p3	
iv. [Kr] 5s2 4d1	
v. [Ne] 3s2 3p3	

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