

Gases Assessment Review

Name _____

Skill - Student will accurately describe the relationship between pressure, volume, temperature and moles.

1. For each of the following scenarios, describe the change in the **VOLUME** of the gas. Label each scenario as exhibiting a direct or inverse relationship.

- Decrease in pressure on the gas
- Decrease in amount (mole) of gas
- Decrease in temperature of gas

Skill - Student will accurately calculate a 2 variable gas law problem.

2. A sample of carbon dioxide gas that has a volume of 5.2 L @ 25°C is heated to 32°C at a constant pressure. Calculate the new volume. **What equation must you use?** Show your work.

Skill - Student will accurately calculate a 3 variable gas law problem (for STP).

3. A sample of Kr gas with a volume of 4.0 L at a temp of -13°C and a pressure of 0.945 atm is heated to 56°C and a pressure of 1.18 atm. Calculate the final volume. **What equation must you use?** Show your work.

4. Hydrogen gas, starting @ 27 °C, at a pressure of 1.25 atm and a volume of 3.2 L is cooled to 2 °C, resulting in a volume of 2.5 L. What is the final pressure of H₂ in the system? Show your work.

Skill - Student will accurately calculate using the ideal gas law.

5. How many grams of chlorine gas are contained in 790.0 mL at 41.0 °C and 750.0 torr pressure? **What equation must you use?** Show your work.

Continue on back →

6. A 2.0L flask contains a mixture of nitrogen gas and oxygen gas at 35°C. The pressure from oxygen gas is 0.95 atm, and the mixture is known to contain 0.050mol nitrogen gas.

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Proceedings of the Marine Safety Council, 1986

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