

Unit 7 Gas Laws Test Review - Answer Key

1. A 0.4 L sample of gas is collected at a temperature of 45°C. At what temperature will the gas fill a 0.725 L container?

LIST GIVENS	FORMULA	GIVENS PLUGGED INTO FORMULA:
P₁ = V ₁ = 0.4 L T ₁ = 318 K P₂ = V ₂ = 0.725 L T ₂ = ?	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{.4}{318} = \frac{.725}{T_2}$ $(.725 \times 318) / 0.4 = T_2$
answer with correct units: T₂ = 576.38 K		

2. If 3 moles of a gas at a pressure of 3.5 atm has a volume of 110 L, what is the temperature?

LIST GIVENS	FORMULA	GIVENS PLUGGED INTO FORMULA:
P = 3.5 atm V = 110 L n = 3 mol R = 0.0821 T = ?	$PV = nRT$	$(3.5 \times 110) = (3 \times 0.0821 \times T)$ $385 / .2463 = T$
answer with correct units: T = 1563.13 K		

3. Identify if the following would cause an INCREASE or DECREASE in pressure in a gaseous system.

- Increase in the size of the container = **pressure decreases**
- More of the same gas is added to the container = **pressure increases**
- Increase in temperature = **pressure increases**
- A different gas is added to the container = **pressure increases**

4. A sealed container of Oxygen gas has a volume of 356 L. If heated from a temperature of 293-K to 318-K, what will the pressure be if the original pressure was 1.2 atm?

LIST GIVENS	FORMULA	GIVENS PLUGGED INTO FORMULA:
P₁ = 1.2 atm V = T ₁ = 293 K P ₂ = ? V = T ₂ = 318 K	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{1.2}{293} = \frac{P_2}{318}$ $(1.2 \times 318) / 293 = P_2$

Gas Laws Test Review Answers

James E. Brady, Fred Senese



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Table of Contents Gas Laws Test Review Answers

1. Understanding the eBook Gas Laws Test Review Answers
 - The Rise of Digital Reading Gas Laws Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Test Review Answers
 - Personalized Recommendations
 - Gas Laws Test Review Answers User Reviews and Ratings
 - Gas Laws Test Review Answers and Bestseller Lists

5. Accessing Gas Laws Test Review Answers Free and Paid eBooks
 - Gas Laws Test Review Answers Public Domain eBooks
 - Gas Laws Test Review Answers eBook Subscription Services
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6. Navigating Gas Laws Test Review Answers eBook Formats
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 - Gas Laws Test Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Test Review Answers
 - Highlighting and Note-Taking Gas Laws Test Review Answers
 - Interactive Elements Gas Laws Test Review Answers
8. Staying Engaged with Gas Laws Test Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Test Review Answers
9. Balancing eBooks and Physical Books Gas Laws Test Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Test Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Test Review Answers
 - Setting Reading Goals Gas Laws Test Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Test Review Answers
 - Fact-Checking eBook Content of Gas Laws Test Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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