

Student Name: (Last, First, Middle Initial) _____, Section: _____, Professor's Name: _____, Date: _____

1. Dry ice (solid carbon dioxide) is also called "snakes." 1.000 grams of dry ice is placed in a 0.050 L container (that is evacuated) at 300.1 K. What is the pressure in the container after all of the dry ice has sublimated?

$$PV = nRT$$

$$\frac{1.000 \text{ g CO}_2}{44} + \frac{1 \text{ mole CO}_2}{44.009 \text{ g CO}_2} = 0.022708 \text{ moles CO}_2$$

$$P = ?$$

$$V = 0.050 \text{ L (1 significant figure)}$$

$$n = 0.022708 \text{ moles (5 significant figures)}$$

$$R = 0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K (5 significant figures)}$$

$$T = 300.1 \text{ K} \rightarrow 300.1 \text{ K (4 significant figures)}$$

$$P = \frac{nRT}{V} = \frac{(0.022708 \text{ moles}) (0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}) (300.1 \text{ K})}{0.050 \text{ L}} = 1.17 \text{ atm}$$

22 grams of a substance B, from the same source

- a) 1.000 atm
- b) 1.000 atm\cdothr
- c) 0.05 atm

2. A sample of chlorine gas is collected over a 0.050 L bottle at atmospheric temperature of pressure. How many moles of chlorine gas are in the container? How many grams?

$$0.0500 \text{ L sample} \rightarrow 0.05 \text{ L}$$

$$\text{Molar Mass of Chlorine (molecular, or as a substance)} = 70.906 \text{ g/mole}$$

What fraction is that

$$\frac{0.05 \text{ L}}{22.4} + \frac{1 \text{ mole CO}_2}{22.4 \text{ L CO}_2} = 0.0022 \text{ moles CO}_2$$

$$\frac{0.05 \text{ L}}{22.4} + \frac{1 \text{ mole CO}_2}{22.4 \text{ L CO}_2} = \frac{0.0022 \text{ L CO}_2}{0.0022 \text{ L CO}_2} = 0.0022 \text{ g CO}_2$$

$$0.05 \text{ L (one sig fig) gas over } PV = nRT$$

$$P = ? \text{ atm (1 significant figure)}$$

$$V = 0.050 \text{ L (1 significant figure)}$$

$$R = 0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K (5 significant figures)}$$

$$T = 273 \text{ K (1 significant figure)}$$

Gas Laws Extra Practice

A Loxley



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