

Gas Stoichiometry problems

1. Calculate the mass of hydrogen peroxide needed to obtain 0.460 L of oxygen gas at STP.



$$.460\text{L O}_2 \left| \frac{1\text{ mole O}_2}{22.4\text{L O}_2} \right| \left| \frac{2\text{ mole H}_2\text{O}_2}{1\text{ mole O}_2} \right| \left| \frac{34.0\text{g H}_2\text{O}_2}{1\text{ mole H}_2\text{O}_2} \right| = \boxed{1.40\text{g}}$$

Given:

$$\begin{aligned} V &= .460\text{L} \\ T &= 273\text{K} \\ P &= 1\text{atm} \\ n &= ? \end{aligned}$$

2. Sodium azide (NaN_3) is used in some air bags. The impact of a collision causes the sodium azide to decompose and form solid Sodium and nitrogen gas. Calculate the volume of nitrogen gas generated at 80.0°C and 823 mm Hg by the decomposition of 60.0 g of NaN_3 .

Balanced equation:



Given:

$$\begin{aligned} T &= 353\text{K} \\ P &= 823\text{ mm Hg} \\ V_{\text{N}_2} &= ? \end{aligned}$$

$$823\text{ mm Hg} \left| \frac{1\text{ atm}}{760\text{ mm Hg}} \right| = 1.08\text{ atm}$$

$$60.0\text{g NaN}_3 \left| \frac{1\text{ mole NaN}_3}{65.0\text{g NaN}_3} \right| \left| \frac{3\text{ moles N}_2}{2\text{ moles NaN}_3} \right| = 1.38\text{ moles N}_2$$

$$V_{\text{N}_2} = \frac{nRT}{P} = \frac{(1.38\text{ moles}) \left(\frac{0.0821\text{ L atm}}{\text{mole K}} \right) (353\text{K})}{1.08\text{ atm}} = 37.0\text{L}$$

3. When heated, solid potassium chlorate (KClO_3) decomposes to form solid potassium chloride and oxygen gas. If 20.8 g of potassium chlorate decomposes, how many liters of oxygen gas will form at STP?

Balanced equation:



Given:

$$\begin{aligned} m &= 20.8\text{g} \\ T &= 273\text{K} \\ P &= 1\text{atm} \\ V &= ? \end{aligned}$$

STP

$$20.8\text{g KClO}_3 \left| \frac{1\text{ mole KClO}_3}{122.55\text{g KClO}_3} \right| \left| \frac{3\text{ moles O}_2}{2\text{ moles KClO}_3} \right| = .255\text{ moles O}_2$$

$$V_{\text{O}_2} = \frac{nRT}{P} = \frac{(0.255\text{ moles}) \left(\frac{0.0821\text{ L atm}}{\text{mole K}} \right) (273\text{K})}{1\text{ atm}}$$

$$\boxed{V_{\text{O}_2} = 5.72\text{ Liters}}$$

OR - Since at STP - do in one step.

$$20.8\text{g KClO}_3 \left| \frac{1\text{ mole KClO}_3}{122.55\text{g KClO}_3} \right| \left| \frac{3\text{ moles O}_2}{2\text{ moles KClO}_3} \right| \left| \frac{22.4\text{L O}_2}{1\text{ mole O}_2} \right| = \boxed{5.70\text{L}}$$

Gas Stoichiometry Worksheet Answers

Sydney Young

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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