

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$$PV = nRT$$

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$$

$$2.34 \text{ L Cl}_2$$



- a. Balance the above reaction.

- b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$$n = x$$

$$V = 0.628 \text{ L}$$

$$T_1 = 25 + 273 = 298 \text{ K}$$

$$P_1 = 1.05 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.628 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (298 \text{ K})}{1.05 \text{ atm}} = 14.6 \text{ L}$$

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.

- b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$$V = \frac{nRT}{P}$$

$$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mole}}{158 \text{ g}} \right| \left| \frac{1 \text{ mole Cl}_2}{2 \text{ moles KMnO}_4} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mole Cl}_2} \right| = 0.709 \text{ L Cl}_2$$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.

- b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$$PV = nRT$$

$$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mole}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{1 \text{ mole}} \right| = 0.204 \text{ mol}$$

$$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mole}}{18 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{3 \text{ moles H}_2\text{O}} \right| = 0.381 \text{ mol}$$

$$T_1 = 24 + 273 = 297 \text{ K}$$

$$P_1 = 752 / 760 = 0.989 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.204 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (297 \text{ K})}{0.989 \text{ atm}} = 5.03 \text{ L}$$

Gas Stoichiometry Sheet Answers

Lei Shi

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Table of Contents Gas Stoichiometry Sheet Answers

1. Understanding the eBook Gas Stoichiometry Sheet Answers
 - The Rise of Digital Reading Gas Stoichiometry Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Stoichiometry Sheet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Gas Stoichiometry Sheet Answers
 - Personalized Recommendations
 - Gas Stoichiometry Sheet Answers User Reviews and Ratings
 - Gas Stoichiometry Sheet Answers and Bestseller Lists
5. Accessing Gas Stoichiometry Sheet Answers Free and Paid eBooks
 - Gas Stoichiometry Sheet Answers Public Domain eBooks
 - Gas Stoichiometry Sheet Answers eBook Subscription Services
 - Gas Stoichiometry Sheet Answers Budget-Friendly Options
6. Navigating Gas Stoichiometry Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Stoichiometry Sheet Answers Compatibility with Devices
 - Gas Stoichiometry Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Stoichiometry Sheet Answers
 - Highlighting and Note-Taking Gas Stoichiometry Sheet Answers
 - Interactive Elements Gas Stoichiometry Sheet Answers
8. Staying Engaged with Gas Stoichiometry Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Stoichiometry Sheet Answers
9. Balancing eBooks and Physical Books Gas Stoichiometry Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Stoichiometry Sheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Stoichiometry Sheet Answers
 - Setting Reading Goals Gas Stoichiometry Sheet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Gas Stoichiometry Sheet Answers
 - Fact-Checking eBook Content of Gas Stoichiometry Sheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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