

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Introduction

Standard Conditions

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760

22.4 L = 1 mol at STP

• STP (standard temperature and pressure) is 0°C and 1 atm.

• The gas laws:

Gas Law	Relationship
Boyle's Law	$P \propto \frac{1}{V}$
Charles's Law	$V \propto T$
Gay-Lussac's Law	$P \propto T$
Avogadro's Law	$V \propto n$
Combined Gas Law	$\frac{PV}{nT} = k$
Ideal Gas Law	$PV = nRT$

• The constant R used in the ideal gas law depends on the unit of pressure:

- a. $0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$
- b. $62.36 \text{ L}\cdot\text{mmHg}/\text{mol}\cdot\text{K}$

Notes:

- Temperature must always be in Kelvin ($K = C + 273$).
- If you are given two different units of pressure, convert first.
- Create two columns below (initial and final). Write what you're given in the problem. Mark the two that use the variable you're given, and solve for the unknown. For ideal gas laws, write what you are given, insert those in equation and solve for unknown.
- Avogadro's Law: Solve for n (moles) to use ideal gas law.
- Graham's Law: You will need to know the molecular mass for these problems. Double and triple check you have the right of effusion and molecular masses in the correct units before solving the equation (this is an $\text{m}^{-1/2}$ relationship). You must evaluate the square root or square both sides of the equation before moving things.

Charles' Law $V \propto T$

1. The temperature of a gas is directly proportional to its volume. If the temperature of a gas is doubled, the volume will also double.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the volume at 127°C and 1.0 atm?

Gay-Lussac's Law $P \propto T$

1. The pressure of a gas is directly proportional to its temperature. If the temperature of a gas is doubled, the pressure will also double.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the pressure at 127°C and 1.0 atm?

Boyle's Law $P \propto \frac{1}{V}$

1. The pressure of a gas is inversely proportional to its volume. If the volume of a gas is doubled, the pressure will be halved.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the pressure at 127°C and 1.0 atm?

Avogadro's Law $V \propto n$

1. The volume of a gas is directly proportional to the number of moles. If the number of moles of a gas is doubled, the volume will also double.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the volume at 127°C and 1.0 atm?

Combined Gas Law $\frac{PV}{nT} = k$

1. The combined gas law relates the pressure, volume, and temperature of a gas. If the pressure of a gas is doubled, the volume will be halved, and the temperature will be doubled.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the volume at 127°C and 1.0 atm?

Ideal Gas Law $PV = nRT$

1. The ideal gas law relates the pressure, volume, number of moles, and temperature of a gas. If the pressure of a gas is doubled, the volume will be halved, and the temperature will be doubled.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the volume at 127°C and 1.0 atm?

All Gas Laws Included!

Gas Laws Packet Answers

David Kirk



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