

**12 • The Gas Laws****THE COMBINED GAS LAW**

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{or} \quad P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$$

$$K = ^\circ C + 273$$

*See  
Scratch  
paper*

Complete the following chart.

|   | $P_1$     | $V_1$   | $T_1$                      | $P_2$     | $V_2$   | $T_2$                          |
|---|-----------|---------|----------------------------|-----------|---------|--------------------------------|
| 1 | 1.50 atm  | 3.00 L  | <del>20.0°C</del><br>293 K | 2.50 atm  |         | <del>30.0°C</del><br>303 K     |
| 2 | 720. torr | 256. mL | <del>25.0°C</del><br>298 K |           | 250. mL | <del>50.0°C</del><br>323 K     |
| 3 | 600. mmHg | 2.50 L  | <del>22.0°C</del><br>325 K | 760. mmHg | 1.80 L  |                                |
| 4 |           | 750. mL | <del>0.00°C</del><br>273 K | 2.00 atm  | 500. mL | <del>25.0°C</del><br>298 K     |
| 5 | 95.0 kPa  | 4.00 L  |                            | 101. kPa  | 6.00 L  | <del>498.°C</del><br>471. K or |
| 6 | 650. torr |         | <del>100.°C</del><br>373 K | 900. torr | 225. mL | <del>150.°C</del><br>423 K     |
| 7 | 850. mmHg | 1.50 L  | <del>15.0°C</del><br>288 K |           | 2.50 L  | <del>30.0°C</del><br>303 K     |
| 8 | 125. kPa  | 125. mL |                            | 100. kPa  | 100 mL  | <del>75.0°C</del><br>348 K     |

# Gas Laws 3 Answer Key

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