

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol,  $C_2H_5OH$ , are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL } C_2H_5OH}{1000 \text{ mL } C_2H_5OH} \times \frac{1 \text{ L } C_2H_5OH}{1 \text{ L } C_2H_5OH} \times \frac{1.5 \text{ mole } C_2H_5OH}{1 \text{ L } C_2H_5OH} = 0.098 \text{ moles } C_2H_5OH$$

2. How many liters of a 6.0 M solution of acetic acid  $CH_3COOH$ , contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles } CH_3COOH}{6.0 \text{ moles } CH_3COOH} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M  $H_2SO_4$  = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M  $Ca(OH)_2$  = 1 N

# Ghs Honors Chemistry The Molarity Worksheet Answers

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## **Table of Contents Ghs Honors Chemistry The Molarity Worksheet Answers**

1. Understanding the eBook Ghs Honors Chemistry The Molarity Worksheet Answers
  - The Rise of Digital Reading Ghs Honors Chemistry The Molarity Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ghs Honors Chemistry The Molarity Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ghs Honors Chemistry The Molarity Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ghs Honors Chemistry The Molarity Worksheet Answers
  - Personalized Recommendations
  - Ghs Honors Chemistry The Molarity Worksheet Answers User Reviews and Ratings
  - Ghs Honors Chemistry The Molarity Worksheet Answers and Bestseller Lists
5. Accessing Ghs Honors Chemistry The Molarity Worksheet Answers Free and Paid eBooks
  - Ghs Honors Chemistry The Molarity Worksheet Answers Public Domain eBooks
  - Ghs Honors Chemistry The Molarity Worksheet Answers eBook Subscription Services
  - Ghs Honors Chemistry The Molarity Worksheet Answers Budget-Friendly Options
6. Navigating Ghs Honors Chemistry The Molarity Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ghs Honors Chemistry The Molarity Worksheet Answers Compatibility with Devices
  - Ghs Honors Chemistry The Molarity Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ghs Honors Chemistry The Molarity Worksheet Answers
  - Highlighting and Note-Taking Ghs Honors Chemistry The Molarity Worksheet Answers
  - Interactive Elements Ghs Honors Chemistry The Molarity Worksheet Answers
8. Staying Engaged with Ghs Honors Chemistry The Molarity Worksheet Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ghs Honors Chemistry The Molarity Worksheet Answers
9. Balancing eBooks and Physical Books Ghs Honors Chemistry The Molarity Worksheet Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ghs Honors Chemistry The Molarity Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ghs Honors Chemistry The Molarity Worksheet Answers
- Setting Reading Goals Ghs Honors Chemistry The Molarity Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ghs Honors Chemistry The Molarity Worksheet Answers
- Fact-Checking eBook Content of Ghs Honors Chemistry The Molarity Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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