

## CHAPTER 21 REVIEW

***Nuclear Chemistry*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ Based on the information about the three elementary particles in the text, which has the greatest mass?
  - (a) the proton
  - (b) the neutron
  - (c) the electron
  - (d) They all have the same mass.
2. \_\_\_\_\_ The force that keeps nucleons together is
  - (a) a strong nuclear force.
  - (b) a weak nuclear force.
  - (c) an electromagnetic force.
  - (d) a gravitational force.
3. \_\_\_\_\_ The stability of a nucleus is most affected by the
  - (a) number of neutrons.
  - (b) number of protons.
  - (c) number of electrons.
  - (d) ratio of neutrons to protons.
4. \_\_\_\_\_ If an atom should form from its constituent particles,
  - (a) matter is lost and energy is taken in.
  - (b) matter is lost and energy is released.
  - (c) matter is gained and energy is taken in.
  - (d) matter is gained and energy is released.
5. \_\_\_\_\_ For atoms of a given mass number, those with greater mass defects, have
  - (a) smaller binding energies per nucleon.
  - (b) greater binding energies per nucleon.
  - (c) the same binding energies per nucleon as those with smaller mass defects.
  - (d) variable binding energies per nucleon.
6. Based on **Figure 1.1** of the text, which isotope of He, helium-3 or helium-4,
  - \_\_\_\_\_ a. has the smaller binding energy per nucleon?
  - \_\_\_\_\_ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is \_\_\_\_\_.
8. Nuclides of the same element have the same \_\_\_\_\_.

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**John S. Phillips, Cheryl Wistrom**



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