Name:	Ashley Maddison	Date:	11/28/2020	

#### Student Exploration: Ionic Bonds

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

 Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

Nate can give Clara 1 marker, so they each have a full set.

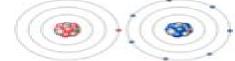
Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

Maggie has to give away 2 markers to Fred and Florence so each has 8.

#### Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *lonic Bonds* Gizmo allows you to explore how **lonic bonds** form.

To begin, check that **Sodium (Na)** and **Chlorine (CI)** are selected from the menus at right. Click **Play** ( b) to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



Each atom consists of a central nucleus and several shells that contain electrons. The outermost electrons
are called valence electrons. (Inner electrons are not shown.)

How many valence electrons does each atom have? Sodium:

n: 1 Chlorine: 7

- Click Pause ( .... ). Elements can be classified as metals and nonmetals. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. Electron affinity is a measure of how tightly the valence electrons are held.
  - A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal?
     Sodium
     Which is a nonmetal? Chlorine
  - B. Try moving an electron from the metal to the nonmetal. What happens?

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#### Table of Contents Gizmo Answers Ionic Bonds

- 1. Understanding the eBook Gizmo Answers Ionic Bonds
  - The Rise of Digital Reading Gizmo Answers Ionic Bonds
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Gizmo Answers Ionic Bonds
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gizmo Answers Ionic Bonds
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Gizmo Answers Ionic Bonds
  - Personalized Recommendations
  - Gizmo Answers Ionic Bonds User Reviews and Ratings
  - Gizmo Answers Ionic Bonds and Bestseller Lists
- 5. Accessing Gizmo Answers Ionic Bonds Free and Paid eBooks
  - o Gizmo Answers Ionic Bonds Public Domain eBooks
  - o Gizmo Answers Ionic Bonds eBook Subscription Services

- Gizmo Answers Ionic Bonds Budget-Friendly Options
- 6. Navigating Gizmo Answers Ionic Bonds eBook Formats
  - o ePub, PDF, MOBI, and More
  - Gizmo Answers Ionic Bonds Compatibility with Devices
  - Gizmo Answers Ionic Bonds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gizmo Answers Ionic Bonds
  - Highlighting and Note-Taking Gizmo Answers Ionic Bonds
  - Interactive Elements Gizmo Answers Ionic Bonds
- 8. Staying Engaged with Gizmo Answers Ionic Bonds
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gizmo Answers Ionic Bonds
- 9. Balancing eBooks and Physical Books Gizmo Answers Ionic Bonds
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gizmo Answers Ionic Bonds
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gizmo Answers Ionic Bonds
  - Setting Reading Goals Gizmo Answers Ionic Bonds
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gizmo Answers Ionic Bonds
  - Fact-Checking eBook Content of Gizmo Answers Ionic Bonds
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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