

1/5/16

Mole Ratio Worksheet

Key

① Given this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, write the following molar ratios:

a) N_2 / H_2 1:3

b) N_2 / NH_3 1:2

c) H_2 / NH_3 3:2

② Given the following equation: $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$, write the following molar ratios:

a) $\text{H}_2 / \text{H}_2\text{S}$ 8:8 $\rightarrow$ 1:1 (reduced)

b) H_2 / S_8 8:1

c) $\text{H}_2\text{S} / \text{S}_8$ 8:1

③ Answer the following questions for this equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

a) What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio? 2:2 $\rightarrow$ 1:1 (reduced)

b) Suppose you had 20 moles of H_2 on hand and plenty of O_2 , how many moles of H_2O could you make? $\frac{20 \text{ mole H}_2}{2 \text{ H}_2} \times \frac{2 \text{ H}_2\text{O}}{1} = \boxed{20 \text{ mole H}_2\text{O}}$

c) What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio?

1:2

d) Suppose you had 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make? $\frac{20 \text{ mol O}_2}{1 \text{ O}_2} \times \frac{2 \text{ H}_2\text{O}}{1} = \boxed{40 \text{ mol H}_2\text{O}}$

4) Use this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, for the following problems

a) If you used 1 mole of N_2 , how many moles of NH_3 could be produced? $\boxed{2}$

b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required? $\frac{10 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{1 \text{ N}_2}{1} = \boxed{5 \text{ mol N}_2}$

c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made? $\frac{3 \text{ mol H}_2}{3 \text{ H}_2} \times \frac{2 \text{ NH}_3}{1} = \boxed{2 \text{ mol NH}_3}$

d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required? $\frac{0.6 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{3 \text{ H}_2}{1} = \boxed{0.9 \text{ mol H}_2}$

Finding Mass Using Mole Ratio Answers

Rachel S Tattersall



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Table of Contents Finding Mass Using Mole Ratio Answers

1. Understanding the eBook Finding Mass Using Mole Ratio Answers
 - The Rise of Digital Reading Finding Mass Using Mole Ratio Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Finding Mass Using Mole Ratio Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Finding Mass Using Mole Ratio Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Finding Mass Using Mole Ratio Answers
 - Personalized Recommendations
 - Finding Mass Using Mole Ratio Answers User Reviews and Ratings
 - Finding Mass Using Mole Ratio Answers and Bestseller Lists
5. Accessing Finding Mass Using Mole Ratio Answers Free and Paid eBooks
 - Finding Mass Using Mole Ratio Answers Public Domain eBooks
 - Finding Mass Using Mole Ratio Answers eBook Subscription Services
 - Finding Mass Using Mole Ratio Answers Budget-Friendly Options
6. Navigating Finding Mass Using Mole Ratio Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Finding Mass Using Mole Ratio Answers Compatibility with Devices
 - Finding Mass Using Mole Ratio Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Finding Mass Using Mole Ratio Answers
 - Highlighting and Note-Taking Finding Mass Using Mole Ratio Answers
 - Interactive Elements Finding Mass Using Mole Ratio Answers
8. Staying Engaged with Finding Mass Using Mole Ratio Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Finding Mass Using Mole Ratio Answers
- 9. Balancing eBooks and Physical Books Finding Mass Using Mole Ratio Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Finding Mass Using Mole Ratio Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Finding Mass Using Mole Ratio Answers
 - Setting Reading Goals Finding Mass Using Mole Ratio Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Finding Mass Using Mole Ratio Answers
 - Fact-Checking eBook Content of Finding Mass Using Mole Ratio Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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