

Chapter 1: Stoichiometry

Q4.

Sol: $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$

Since 1 mole of CaCO_3 produces 1 mole of CaO , So

1 mole of CaO will produce 1 mole of CaCO_3 .

\therefore 1 mole of CaO \longrightarrow 1 mole of CaCO_3 .

Sol: $\text{BaCl}_2 \longrightarrow \text{BaSO}_4$

Given:

5 moles of BaCl_2 contains 5 moles of Ba & 10 moles of Cl & each form 1 mole of BaSO_4 , \therefore since 1 mole of BaSO_4 contains 1 mole of Ba .

Sol: $2 \text{Na} \longrightarrow \text{Na}_2\text{O}$

5 moles O_2 given:

\therefore 2 moles produces 1 mole of Na_2O .

\therefore 5 \longrightarrow $\frac{5}{2}$

\therefore 5 \longrightarrow $5 \times \frac{1}{2} = 2.5$ moles of Na_2O .

Q1.

Ans: $\text{Ca} + \frac{1}{2} \text{O}_2 \longrightarrow \text{CaO}$ (balanced chemical equation)

Wt. of $\text{Ca} = 40 \text{ g/mole}$

So, moles of $\text{Ca} = \frac{80}{40} = 2 \text{ moles}$

From above balanced chemical equation: 1 moles of Ca reacts with $\frac{1}{2}$ moles of O_2 .

Wt. of O_2 with which 80g of Ca react $= \frac{1}{2} \times 32 = 16 \text{ g}$

Ans.

Q2.

Sol: $\text{NaCl} \longrightarrow \text{NaClO}_3$

58 gms.

If we have not balanced the chemical equation, then we have to apply PMOL rule as per equivalent ratio.

By PMOL rule Cl atom

From NaCl \longrightarrow NaClO_3 we have moles \therefore 1 mol. of Cl identification system.

So, 1 mole of NaCl \rightarrow NaClO_3 \therefore 2 = 1 = 1 mole of NaCl \rightarrow NaClO_3 .

$\frac{58}{58.5} = 1$ mole of NaCl .

moles of $\text{O}_2 = \frac{48}{32}$

Modern Approach To Chemical Calculations Solutions

JS Bruner



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