Deather

Chemistry - Unit 5 Worksheet 2

 An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

 $3.3 \text{ g} \times \underline{1 \text{ mole}} = 5.2 \times 10^{\circ} \text{ mol or } 0.052 \text{ mol}$ 63.5 g

Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

4.42 g = 1 mole = 7.92 x 10° mol or 0.0792 mol 55.8 g

A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

0.050 g 1 mole = 2.78 x 10° mol → 2.8 x 10° mol

4. What mass of water would you need to have 15.0 moles of H2O?

15.0 moles 18.0 g = 2.70 x 10² g

 One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

737 g 1 mole = 12.6 mol

A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

0.050 moles 84.0 g = 4.2 g

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2 50 kg of rust?

Fe 111.6

48.0

2.50 kg 1000 g 1 mole = 15.7 mol Fe₂O₃ 1kg 159.6 g CI 35.5

58.5 a

No. 23.0

12.0

48.0

84.0

Modeling Chemistry Unit 5 Test Answers

Ying Liu

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