


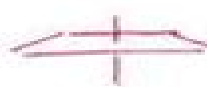

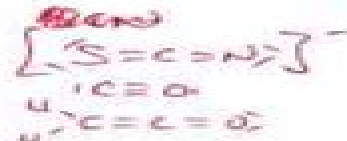



10 • Orbital Hybridization & Molecular Orbitals

PRACTICE TEST

- Which hybridization is associated with a steric number of 3?
a) sp SN=2
b) sp^2 3
c) sp^3 4
- The molecule BrF_3 has a steric number of 5 on the central atom?
a) 3 b) 4 **c) 5** d) 6

- What is the hybridization of Br in BrF_3 ?
a) sp
b) sp^2
c) sp^3
d) sp^3d SN=5
e) sp^3d^2
- How many equivalent sp^3d orbitals are there?
a) 3 b) 1 **c) 5** d) 6
- What type of hybridization is associated with a square planar molecular shape?
a) sp^3
b) sp^2
c) sp
d) sp^3d^2 c)

- What shape for electron pairs is associated with sp^3d^2 hybridization?
a) linear d) tetrahedral
b) square planar **c) octahedral**
c) bent
- What hybridization is predicted for phosphorus in the PCl_5 molecule?
a) sp^2
b) sp^3
c) sp
d) sp^3d^2

tetrahedral shape = pyramidal
- A double bond contains 1 sigma bond(s) and 1 pi bond(s).
a) 0, 2 b) 1, 2 c) 2, 0 **d) 1, 1**
- What angle exists between orbitals in sp^3d^2 hybrid orbitals?
a) 90.0° d) 120.0°
b) 180.0° e) 78.5° c) 109.5°

- Which of the following elements is most likely to display sp^3d hybridization? SN=5
a) oxygen d) carbon
b) nitrogen e) boron } no d orbitals
c) phosphorus
- How many sigma (σ) and pi (π) electron pairs are in a carbon dioxide molecule? $O=C=O$
a) four σ and zero π d) two σ and four π
b) three σ and two π e) one σ and three π
c) two σ and two π
- What is the hybridization of the oxygen atoms in CH_3OH and CO_2 , respectively? $O=C=O$
a) sp^3 , sp^3 d) sp^2 , sp^2 SN=3
b) sp^3 , sp^2 e) sp^3 , sp sp^2
c) sp^2 , sp^2

- All of the following species contain two π -bonds EXCEPT
a) SCN^- 16e- d) OCS 16e-
b) CO 10e- **e) NO^+** 12e-
c) H_2CCO 16e-



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