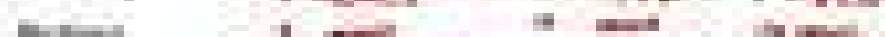
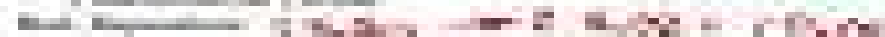


Stoichiometry Worksheet 2: Percent Yield

What parts of a stoichiometry problem:

- Identify a balanced chemical equation.
- Identify which mass of a reactant or product is given and which mass of a reactant or product is needed.
- Identify any other data needed. (Chemical steps, given mass, molar masses, etc.)

Using the following example as a guide, identify the given and needed masses of a reactant or product. What mass of H_2O is produced from 10.0 g of H_2 and 10.0 g of O_2 ?



$$\text{Mass } \text{H}_2\text{O} = \frac{\text{mass } \text{H}_2\text{O}}{\text{mass } \text{H}_2} \times \text{mass } \text{H}_2$$

$$1.00 \text{ mol } \text{H}_2 = \frac{2.016 \text{ g } \text{H}_2}{2.016 \text{ g } \text{H}_2} \times 10.0 \text{ g } \text{H}_2$$

$$1.00 \text{ mol } \text{H}_2 = \frac{2.016 \text{ g } \text{H}_2}{2.016 \text{ g } \text{H}_2} \times 10.0 \text{ g } \text{H}_2$$

Identify which mass of a reactant or product is given and which mass of a reactant or product is needed. What mass of H_2O is produced from 10.0 g of H_2 and 10.0 g of O_2 ?



$$\begin{aligned} 1.00 \text{ mol } \text{H}_2 &= \frac{2.016 \text{ g } \text{H}_2}{2.016 \text{ g } \text{H}_2} \times 10.0 \text{ g } \text{H}_2 \\ 1.00 \text{ mol } \text{O}_2 &= \frac{32.0 \text{ g } \text{O}_2}{32.0 \text{ g } \text{O}_2} \times 10.0 \text{ g } \text{O}_2 \end{aligned}$$

What mass of H_2O is produced from 10.0 g of H_2 and 10.0 g of O_2 ? (10.0 g of H_2 and 10.0 g of O_2 are given and 10.0 g of H_2O is needed.)



$$1.00 \text{ mol } \text{H}_2 = \frac{2.016 \text{ g } \text{H}_2}{2.016 \text{ g } \text{H}_2} \times 10.0 \text{ g } \text{H}_2$$

$$\text{Mass } \text{H}_2\text{O} = \frac{\text{mass } \text{H}_2\text{O}}{\text{mass } \text{H}_2} \times \text{mass } \text{H}_2 = \frac{36.0 \text{ g } \text{H}_2\text{O}}{4.00 \text{ g } \text{H}_2} \times 10.0 \text{ g } \text{H}_2 = 9.00 \text{ g } \text{H}_2\text{O}$$

Identify which mass of a reactant or product is given and which mass of a reactant or product is needed. What mass of H_2O is produced from 10.0 g of H_2 and 10.0 g of O_2 ?



$$10.0 \text{ g } \text{H}_2 \times \left(\frac{1 \text{ mol } \text{H}_2}{2.016 \text{ g } \text{H}_2} \right) = 4.96 \text{ mol } \text{H}_2$$

$$\frac{10.0 \text{ g } \text{O}_2}{32.0 \text{ g } \text{O}_2} = 0.3125 \text{ mol } \text{O}_2$$

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Siddappa N.Byrareddy



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