At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br₂] was measured to be 0.13 M. What is K_{eq} for this reaction at this temperature?

	2HBr =	= H2 +	Br2
[x]	0.50	0	0
[C]	-0-26	+0.13	+0.13
IE]	0.24	0.13	0.13

Keg =	IH2][Br]2	= (0.13)2	=0.29
	100001100000000000000000000000000000000	150000000000000	

Answer Keg = 0.29

 When 1.0 mol of NH_{3(g)} and 0.40 mol of N_{2(g)} are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH₃ is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

	aNH3 &	= 3 Na	+ N2
[I]	0.20	0	0.080
EC3	-0.044	+0.066	+0.022
LEJ.	0.156	0.066	0.102

a) Calculate the equilibrium concentrations of all three species.

[NH₃] - 0.16 M [H₂] - 0.066 M [N₂] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H₂ are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

d) How many moles of N₂ are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

Modeling Chemistry U11 Ws 1 Answers

RC Schank

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "Modeling Chemistry U11 Ws 1 Answers," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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