

Name Tara ShettyDate 9/24/15 Pd 5

## Chemistry – Unit 4 Worksheet 3

Use the following information about the masses of elements in each pair of compounds to help you suggest formulas that account for these ratios.

## 1. Compounds of carbon and oxygen

Compound A: 57.1 g O / 42.9 g C

Compound B: 72.7 g O and 27.3 C

$$A = \frac{57.1}{42.9} = 1 \frac{1}{3}$$

$$B = \frac{72.7}{27.3} = 2 \frac{2}{3}$$

a. Determine the value of the ratio  $\frac{\text{mass O}}{\text{mass C}}$  in each compound. A  $1 \frac{1}{3}$  B  $2 \frac{2}{3}$

b. How does the mass ratio for compound B compare to that in compound A?

compound B is larger than compound A

c. Express these ratios as improper fractions.  $\frac{57.1}{42.9} + \frac{72.7}{27.3} \left( \frac{4}{3} + \frac{8}{3} \right)$

d. For each hypothesis, sketch particle diagrams for the compounds of A and B that account for these mass ratios. Write the formula for the compound in each diagram.

## Hypothesis 1

Atoms of C and O have the same mass

A  $\frac{4}{3}$ B  $\frac{8}{3}$ 

## Hypothesis 2

Atoms of O are heavier than C atoms by the ratio in compound A.

A



B



# Modeling Chemistry Unit 10 2 Answer

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