Oxidation and Reduction

What happens when electrons are transferred in a chemical reaction?

Why?

Silver tarnishes when it comes in contact with sulfur compounds in the air. Copper gets coated in beautiful green patina as it ages. Metals rust or corrode in the presence of air and water. Minerals (ionic compounds) found in ore can be decomposed with the use of electricity to produce pure metals and nonmetals. All of these reactions are examples of oxidation and reduction, otherwise known as redox reactions. In this activity you will explore what is happening at the atomic level in redox reactions.

Model 1 - Redox Reactions

Redox Reactions

A.
$$Zn(s) + Cu^{2s}(aq) \rightarrow Zn^{2s}(aq) + Cu(s)$$
 B. $2I^{-}(aq) + S_2O_s^{2-}(aq) \rightarrow I_2(s) + 2SO_s^{2-}(aq)$

C.
$$4Fe(s) + 3O_s(g) \rightarrow 2Fe_sO_s(s)$$
 D. $4H^*(aq) + MnO_s^*(aq) + 3Fe^{3*}(aq) \rightarrow 3Fe^{3*}(aq) + MnO_s(aq) + 2H_sO(l)$

Nonredox Reactions

$$E. \ HCl(g) + H_2O(l) \rightarrow H_3O'(aq) + Cl'(aq) \quad E. \ 2NaOH(aq) + H_2SO_4(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(l)$$

G.
$$Ba^{2s}(aq) + 2OH^{s}(aq) \rightarrow Ba(OH)_{s}(s)$$
 H. $2AgNO_{s}(aq) + CaCl_{s}(aq) \rightarrow Ca(NO_{s})_{s}(aq) + 2AgCl(s)$

- 1. What two types of reactions are shown in Model 1?
- Examine the redox and nonredox reactions in Model 1. Is/are there any feature(s) in the redox reactions that would allow you to identify them as redox reactions? If yes, use specific examples from Model 1 to support your answer.

Model 1 Redox Reactions Answers Pogil

Thomas Griffiths



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