

Ions that form soluble compounds	Exceptions	Ions that form soluble compounds	Exceptions
Group 1 ions ( $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , ...)	-	Carbonate ( $\text{CO}_3^{2-}$ )	When combined with group 1 ions or ammonium ( $\text{NH}_4^+$ )
Ammonium ( $\text{NH}_4^+$ )	-	Chromate ( $\text{CrO}_4^{2-}$ )	When combined with group 1 ions, $\text{Ca}^{2+}$ , $\text{Mg}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )
Nitrate ( $\text{NO}_3^-$ )	-	Phosphate ( $\text{PO}_4^{3-}$ )	When combined with group 1 ions or ammonium ( $\text{NH}_4^+$ )
Acetate ( $\text{C}_2\text{H}_3\text{O}_2^-$ or $\text{CH}_3\text{COO}^-$ )	-	Sulfide ( $\text{S}^{2-}$ )	When combined with group 1 ions or ammonium ( $\text{NH}_4^+$ )
Hydrogen Carbonate ( $\text{HCO}_3^-$ )	-	Hydroxide ( $\text{OH}^-$ )	When combined with group 1 ions, $\text{Ca}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Sr}^{2+}$ or ammonium ( $\text{NH}_4^+$ )
Chlorate ( $\text{ClO}_3^-$ )	-	Sulfates ( $\text{SO}_4^{2-}$ )	When combined with $\text{Ag}^+$ , $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , and $\text{Pb}^{2+}$
Perchlorate ( $\text{ClO}_4^-$ )	-	Halides ( $\text{Cl}^-$ , $\text{Br}^-$ , $\text{I}^-$ )	When combined with $\text{Ag}^+$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$

Solubility is important because it directly impacts the amount of radioisotope available for the reaction. It may affect distribution in the radiation source resulting in dosimetry issues. For example, in Benega et. al. [34] a phosphorus-32 radioactive source for spinal cancer treatment was developed by mixing the radioisotope with a catalyst solution. This solution is then added to an epoxy resin. Both solutions need to result in a homogenous product. This won't be achievable if the isotope doesn't properly mix with the catalyst solution (water based).

The solubility product constant,  $K_{sp}$ , is the equilibrium constant for a solid substance dissolving in an aqueous solution. It represents the level at which a solute will dissolve in solution. Equation 17 shows the correlation.



# Model 1 Rules Of Solubility In Aqueous Solutions

**Michael Seilmaier**



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## **Table of Contents Model 1 Rules Of Solubility In Aqueous Solutions**

1. Understanding the eBook Model 1 Rules Of Solubility In Aqueous Solutions
  - The Rise of Digital Reading Model 1 Rules Of Solubility In Aqueous Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Model 1 Rules Of Solubility In Aqueous Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Model 1 Rules Of Solubility In Aqueous Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Model 1 Rules Of Solubility In Aqueous Solutions
  - Personalized Recommendations
  - Model 1 Rules Of Solubility In Aqueous Solutions User Reviews and Ratings

- Model 1 Rules Of Solubility In Aqueous Solutions and Bestseller Lists
- 5. Accessing Model 1 Rules Of Solubility In Aqueous Solutions Free and Paid eBooks
  - Model 1 Rules Of Solubility In Aqueous Solutions Public Domain eBooks
  - Model 1 Rules Of Solubility In Aqueous Solutions eBook Subscription Services
  - Model 1 Rules Of Solubility In Aqueous Solutions Budget-Friendly Options
- 6. Navigating Model 1 Rules Of Solubility In Aqueous Solutions eBook Formats
  - ePub, PDF, MOBI, and More
  - Model 1 Rules Of Solubility In Aqueous Solutions Compatibility with Devices
  - Model 1 Rules Of Solubility In Aqueous Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Model 1 Rules Of Solubility In Aqueous Solutions
  - Highlighting and Note-Taking Model 1 Rules Of Solubility In Aqueous Solutions
  - Interactive Elements Model 1 Rules Of Solubility In Aqueous Solutions
- 8. Staying Engaged with Model 1 Rules Of Solubility In Aqueous Solutions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Model 1 Rules Of Solubility In Aqueous Solutions
- 9. Balancing eBooks and Physical Books Model 1 Rules Of Solubility In Aqueous Solutions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Model 1 Rules Of Solubility In Aqueous Solutions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Model 1 Rules Of Solubility In Aqueous Solutions
  - Setting Reading Goals Model 1 Rules Of Solubility In Aqueous Solutions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Model 1 Rules Of Solubility In Aqueous Solutions
  - Fact-Checking eBook Content of Model 1 Rules Of Solubility In Aqueous Solutions
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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