

change of the surrounding water. $\Delta H_{\text{metal}} = -\Delta H_{\text{water}}$ If the water temperature increases, then the metal must have released heat. We can find the enthalpy change of the metal with the formula: $\Delta H_{\text{metal}} = -m_{\text{water}} \cdot C_{\text{water}} \cdot \Delta T_{\text{water}}$

Check-for-understanding question:

1. When 25 mL of water containing 0.85 mol HCl is mixed with 40 mL of water containing .05 mol KOH in a calorimeter (with initial temperature 23 °C), the temperature of the solution increases to 25.5 °C. How many kilojoules (kJ) of heat are released by the acid-base reaction?

$$\begin{aligned}\Delta H_{\text{water}} &= m_{\text{water}} \times C_{\text{water}} \times \Delta T_{\text{water}} \\ &= 65 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 2.5^\circ\text{C} \\ &= 679.25 \text{ J} = .67925 \text{ kJ}\end{aligned}$$

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The logo of the American Petroleum Institute, featuring a red, stylized flame or drop shape.

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TEXTBOOK OF CHEMICAL ENGINEERING THERMODYNAMICS K. V. NARAYANAN, 2013-01-11 Designed as an undergraduate level textbook in Chemical Engineering this student friendly thoroughly class room tested book now in its second edition continues to provide an in depth analysis of chemical engineering thermodynamics The book has been so organized that it gives comprehensive coverage of basic concepts and applications of the laws of thermodynamics in the initial chapters while the later chapters focus at length on important areas of study falling under the realm of chemical thermodynamics The reader is thus introduced to a thorough analysis of the fundamental laws of thermodynamics as well as their applications to practical situations This is followed by a detailed discussion on relationships among thermodynamic properties and an exhaustive treatment on the thermodynamic properties of solutions The role of phase equilibrium thermodynamics in design analysis and operation of chemical separation methods is also deftly dealt with Finally the chemical reaction equilibria are skillfully explained Besides numerous illustrations the book contains over 200 worked examples over 400 exercise problems all with answers and several objective type questions which enable students to gain an in depth understanding of the concepts and theory discussed The book will also be a useful text for students pursuing courses in chemical engineering related branches such as polymer engineering petroleum engineering and safety and environmental engineering New to This Edition More Example Problems and Exercise Questions in each chapter Updated section on Vapour Liquid Equilibrium in Chapter 8 to highlight the significance of equations of state approach GATE Questions up to 2012 with answers **Annual Report of Progress--fundamental Research on Occurrence and Recovery of Petroleum** American Petroleum Institute, 1944 **Physical Chemistry for the Biosciences** Raymond Chang, 2005-02-11 This book is ideal for use in a one semester introductory course in physical chemistry for students of life sciences The author's aim is to emphasize the understanding of physical concepts rather than focus on precise mathematical development or on actual experimental details Subsequently only basic skills of differential and integral calculus are required for understanding the equations The end of chapter problems have both physiochemical and biological applications *Biochemistry* J. Stenesh, 2013-06-29 This text is intended for an introductory course in bio metabolism concludes with photosynthesis The last sec chemistry While such a course draws students from vari tion of the book Part IV TRANSFER OF GENETIC INFOR ous curricula all students are presumed to have had at MATION also opens with an introductory chapter and then least general chemistry and one

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Physical Chemistry for the Life Sciences Peter William Atkins,R. George Ratcliffe,Julio de Paula,Mark Wormald,2023 This text provides a balanced presentation of the concepts of physical chemistry and their applications to biology and biochemistry Written to straddle the worlds of physical chemistry and the life sciences it shows students how the tools of physical chemistry can elucidate biological questions

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