35(**)

42Cm2-

Atomic Structure:

- What do these have in common?
 - A) The same number of protons
 - B) The same number of electrons
 - C) The same number of neutrons
 - D) The same size

The element X occurs natural to the extent of 20.0% ¹²X and 80.0% ¹³X. The atomic mass of X is nearest

- A) 12.2
- B) 12.5
- C) 12.8
- D) 13.0

3. What is the valence electron configuration for the element in Period 5, Group 3A?

- A) $5s^25p^4$
- B) 3s23p5
- C) $3s^23p^4$
- D) 5s25p3

4. What is the maximum number of electrons that can occupy an orbital labeled d_{ss}?

- A) 1
- B) 2
- C) 4

D) 10

5. Which electron transition in a hydrogen atom is associated with the largest emission of energy?

- A) n = 2 to n = 1
- B) n = 1 to n = 2
- C) n = 2 to n = 4
- D) n = 4 to n = 2

Molecular Structure and Bonding:

1. Which molecule has exactly two unshared (lone) pairs of electrons on the central atom?

- A) BF₃
- B) OF₂
- C) NE₃
- D) XeF₂

2. The molecule of the type ML₄ consists of four single bonds and no lone pairs. What structure is it expected to assume?

- A) square planar
- B) trigonal pyramidal
- C) trigonal planar
- D) tetrahedral

Mastering Chemistry Chem 113 Exercise Answers

Chao Zhang

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