

Mark Scheme

Electrolysis – Molten

1. Ions cannot move
2. Anode (Remember PANIC: Positive Anode Negative Is Cathode).
3. Elements
4. Electrolyte
5. OILRIG - Oxidation Is Loss, Reduction Is Gain (of electrons).
6. Oxidation
7. Ca^{2+} would go to the cathode, I^- would go to the anode.
8. $\text{Ca}^{2+} + 2\text{e}^- \rightarrow \text{Ca}$
9. $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
10. $\text{CaCl}_2 \rightarrow \text{Ca}^{2+} + 2\text{Cl}^-$

Electrolysis – Solutions

1. NaCl
2. Na^+ , H^+ , OH^- , Cl^-
3. Hydrogen, chlorine, sodium hydroxide.
4. Negative ions (OH^- , Cl^-)
5. Positive Ions (H^+ , Na^+)
6. Chlorine, $2\text{Cl}^- \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^-$
7. Hydrogen, $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$
8. Na^+ and OH^-
9. Sodium hydroxide
10. Aluminium. Oxygen is released at the anode where it reacts with the graphite to form carbon dioxide. Therefore the anode needs to be replaced often.

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Lingjun Ying

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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