

The Gnu Emacs Lisp Reference Manual: Gnu Emacs Version 18 for Unix Users

Gnu Manual Group

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Gnu Emacs Lisp Manual

Robert J. Chassell



Gnu Emacs Lisp Manual:

GNU Emacs LISP Reference Manual 2/2 Bil Lewis,Dan Laliberte,Richard Stallman,2015-11-06 Most of the GNU Emacs text editor is written in the programming language called Emacs Lisp You can write new code in Emacs Lisp and install it as an extension to the editor However Emacs Lisp is more than a mere extension language it is a full computer programming language in its own right You can use it as you would any other programming language Because Emacs Lisp is designed for use in an editor it has special features for scanning and parsing text as well as features for handling files buffers displays subprocesses and so on Emacs Lisp is closely integrated with the editing facilities thus editing commands are functions that can also conveniently be called from Lisp programs and parameters for customization are ordinary Lisp variables This manual attempts to be a full description of Emacs Lisp For a beginner s introduction to Emacs Lisp see An Introduction to Emacs Lisp Programming by Bob Chassell also published by the Free Software Foundation This manual presumes considerable familiarity with the use of Emacs for editing see The GNU Emacs Manual for this basic information Generally speaking the earlier chapters describe features of Emacs Lisp that have counterparts in many programming languages and later chapters describe features that are peculiar to Emacs Lisp or relate specifically to editing This is the GNU Emacs Lisp Reference Manual corresponding to Emacs version 24.5 As Emacs Lisp became such a big project over the years we had to split this reference manual in two parts that are two separate physical books To keep it consistent with our digital manual the references and page numbers cover both physical books as it were one Therefore please note that you probably want to have both parts

The GNU Emacs Lisp Reference Manual Bil Lewis,Dan LaLiberte,Richard M. Stallman,1998

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow 2\text{Cr}^{3+} \text{aq} + 7\text{H}_2\text{O} \text{l} + 2\text{H}^+ \text{aq}$

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{CuCl}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+\text{aq} + \text{Cl}^-\text{aq} + \text{Na}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{Na}^+\text{aq} + \text{Cl}^-\text{aq} + \text{H}_2\text{O} + \text{CO}_2\text{g}$ the spectator ions are Cl^- and Na^+
 net ionic equation $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2\text{g}$ here is a video on spectator ions
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web the six NO_3^- ions and the six Na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3\text{aq} + \text{KClaq} \rightarrow \text{AgCl s} + \text{KNO}_3\text{aq}$ total ionic $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$ net ionic $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s}$ 2

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{l}$ the spectator ions in KOH aq and HNO_3aq are K^+aq and NO_3^-aq how to identify and list the

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no 3 aq ions stay as na aq ions and no 3 aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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