

Name _____

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Chemistry – Unit 10 Worksheet 3

1. Element X has two natural isotopes. The isotope with a mass number of 6 has a relative abundance of 7.5%. The isotope with a mass number of 7 has a relative abundance of 92.5%. Determine the average molar mass for the element from these figures. What is the true identity and atomic number of element X?
2. The element copper is found to contain the naturally occurring isotopes $^{63}_{29}\text{Cu}$ and $^{65}_{29}\text{Cu}$. The relative abundances are 69.1% and 30.9% respectively. Calculate the average molar mass of copper.
3. Uranium has three isotopes with the following percent abundances: $^{234}_{92}\text{U}$ (0.0058%), $^{235}_{92}\text{U}$ (0.71%), $^{238}_{92}\text{U}$ (99.23%). What do you expect the molar mass of uranium to be in whole numbers? Why?
4. A sample of silver as it occurs in nature is 52.0% of isotope $^{107}_{47}\text{Ag}$ and 48.0% of isotope $^{109}_{47}\text{Ag}$. What is the average molar mass of silver? (Compare your result with the value given in the periodic table).

Modeling Chemistry U7 Ws 2 V2 Answers

Y Pai



Modeling Chemistry U7 Ws 2 V2 Answers:

The book delves into Modeling Chemistry U7 Ws 2 V2 Answers. Modeling Chemistry U7 Ws 2 V2 Answers is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry U7 Ws 2 V2 Answers, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 1: Introduction to Modeling Chemistry U7 Ws 2 V2 Answers
 - Chapter 2: Essential Elements of Modeling Chemistry U7 Ws 2 V2 Answers
 - Chapter 3: Modeling Chemistry U7 Ws 2 V2 Answers in Everyday Life
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 - Chapter 5: Conclusion
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