

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We can calculate the relative abundance of an element by counting the number of atoms of each isotope present.

1. Equal numbers of atoms contain the same number of nucleons as the same physical mass of elements. **Simplest System**
2. Abundance of isotopes of an element can give us their relative abundance.

${}^1_1\text{H}$, ${}^2_1\text{H}$, ${}^3_1\text{H}$, ${}^4_2\text{He}$, ${}^{12}_6\text{C}$, ${}^{13}_6\text{C}$, ${}^{16}_8\text{O}$, ${}^{17}_8\text{O}$, ${}^{18}_8\text{O}$

Example: Calculate the relative abundance of oxygen isotopes. First find the density of oxygen gas in standard temperature and pressure (1.43 g/l) and the relative atomic mass of oxygen gas (32).
 Relative abundance is 99.9% oxygen-16, 0.04% oxygen-17, and 0.03% oxygen-18.

Oxygen	Hydrogen	Relative
32 g/mol	2 g/mol	
1 liter	1 liter	

32
Oxygen

16

2
Hydrogen

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32 g/mol	2 g/mol	
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