

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

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Table of Contents Modeling Chemistry Unit Ws 2 Answers

1. Understanding the eBook Modeling Chemistry Unit Ws 2 Answers
 - The Rise of Digital Reading Modeling Chemistry Unit Ws 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit Ws 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit Ws 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit Ws 2 Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit Ws 2 Answers User Reviews and Ratings
 - Modeling Chemistry Unit Ws 2 Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit Ws 2 Answers Free and Paid eBooks

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 - Modeling Chemistry Unit Ws 2 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit Ws 2 Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit Ws 2 Answers
 - Interactive Elements Modeling Chemistry Unit Ws 2 Answers
- 8. Staying Engaged with Modeling Chemistry Unit Ws 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit Ws 2 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit Ws 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit Ws 2 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit Ws 2 Answers
 - Setting Reading Goals Modeling Chemistry Unit Ws 2 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit Ws 2 Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit Ws 2 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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