

Unit 7: Balancing Chemical Reactions • Worksheet 2

Balance the following equations by inserting the proper coefficients.

1. $\underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{SO}_3$
 2. $\underline{2}\text{CH}_4 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO} + \underline{4}\text{H}_2\text{O}$
 3. $\underline{2}\text{P} + \underline{3}\text{Cl}_2 \rightarrow \underline{2}\text{PCl}_3$
 4. $\underline{2}\text{CO} + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CO}_2$
 5. $\underline{2}\text{CH}_4 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{OH}$
 6. $\underline{2}\text{Li} + \underline{1}\text{Br}_2 \rightarrow \underline{2}\text{LiBr}$
 7. $\underline{2}\text{Al}_2\text{O}_3 \rightarrow \underline{4}\text{Al} + \underline{3}\text{O}_2$
 8. $\underline{2}\text{Na} + \underline{2}\text{H}_2\text{O} \rightarrow \underline{2}\text{NaOH} + \underline{1}\text{H}_2$
 9. $\underline{6}\text{CO}_2 + \underline{6}\text{H}_2\text{O} \rightarrow \underline{1}\text{C}_6\text{H}_{12}\text{O}_6 + \underline{6}\text{O}_2$
 10. $\underline{1}\text{H}_2\text{SO}_4 + \underline{2}\text{NaCl} \rightarrow \underline{2}\text{HCl} + \underline{1}\text{Na}_2\text{SO}_4$
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
11. $\underline{3}\text{H}_2 + \underline{1}\text{SO}_2 \rightarrow \underline{1}\text{H}_2\text{S} + \underline{2}\text{H}_2\text{O}$
 12. $\underline{2}\text{CaCO}_3 + \underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CaSO}_4 + \underline{2}\text{CO}_2$
 13. $\underline{2}\text{AgNO}_3 + \underline{1}\text{CaCl}_2 \rightarrow \underline{2}\text{AgCl} + \underline{1}\text{Ca}(\text{NO}_3)_2$
 14. $\underline{2}\text{HCl} + \underline{1}\text{Ba}(\text{OH})_2 \rightarrow \underline{1}\text{BaCl}_2 + \underline{2}\text{H}_2\text{O}$
 15. $\underline{1}\text{H}_3\text{PO}_4 + \underline{3}\text{NaOH} \rightarrow \underline{1}\text{Na}_3\text{PO}_4 + \underline{3}\text{H}_2\text{O}$
 16. $\underline{1}\text{Pb}(\text{NO}_3)_2 + \underline{2}\text{KI} \rightarrow \underline{1}\text{PbI}_2 + \underline{2}\text{KNO}_3$
 17. $\underline{3}\text{CuO} + \underline{2}\text{NH}_3 \rightarrow \underline{1}\text{N}_2 + \underline{3}\text{Cu} + \underline{3}\text{H}_2\text{O}$
 18. $\underline{1}\text{C}_2\text{H}_5\text{OH} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{3}\text{H}_2\text{O}$
 19. $\underline{2}\text{C}_2\text{H}_6 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{COOH} + \underline{2}\text{H}_2\text{O}$
 20. $\underline{3}\text{NO}_2 + \underline{1}\text{H}_2\text{O} \rightarrow \underline{2}\text{HNO}_3 + \underline{1}\text{NO}$

Modeling Chemistry Worksheet Rearranging Atoms

Answer Key

RJ Alexander



Modeling Chemistry Worksheet Rearranging Atoms Answer Key:

Modeling Chemistry Worksheet Rearranging Atoms Answer Key Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Modeling Chemistry Worksheet Rearranging Atoms Answer Key**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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