Unit 6 - Thermodynamics (7-9%)

Changes in a substance's properties or change into a different substance requires an exchange of energy.

 Endothermic and Easthermic Peaceuses: Explain the relationship between experimental observations and energy changes associated with a chemical or physical transformation.

- Temperature changes in a system indicate energy changes.
- Energy changes in a system can be described as endothermic and exothermic processes such as the heating or cooling of a substance, phase changes, or chemical transformations.
- When a chemical reaction occurs, the energy of the system either decreases (exothermic reaction). increases (endothermic reaction), or remains the same. For exothermic reactions, the energy lost by the reacting species (system) is gained by the surroundings, as heat transfer from or work done by the system. Likewise, for endothermic reactions, the system gains energy from the surroundings by heat transfer to or work done on the system.
- The formation of a solution may be an exothermic or endothermic process, depending on the relative streneths of intermolecular Internetticle interactions before and after the dissolution process

Unit 6 Objectives

- Thermodynamics

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