

| U3 - Quest 2 b - Electron Configurations & Bohr Models |       |       |               |
|--------------------------------------------------------|-------|-------|---------------|
| 4                                                      | 3     | 2     | 1             |
| 15-16 pts                                              | 13-14 | 10-12 | 9 pts or less |

1. What is the full electron configuration for Indium (In, atomic number = 49)?
2. Draw a Bohr model of an Indium-115 ( $^{115}\text{In}$ ) atom at ground state. Indicate the number of the protons, neutrons, total electrons, and the number of electrons located in each energy level. Label the nucleus, and energy levels ( $n = 1, n = 2$ , etc.).

#p+ \_\_\_\_\_

#n+ \_\_\_\_\_

#e- \_\_\_\_\_

[total]

 $n = 1$  \_\_\_\_\_ e- $n = 2$  \_\_\_\_\_ e- $n = 3$  \_\_\_\_\_ e- $n = 4$  \_\_\_\_\_ e- $n = 5$  \_\_\_\_\_ e-

3. Write the condensed (noble gas) configurations for each of the following elements. Then determine the number of valence (outermost) electrons in each element.

| Element      | Atomic Number | Condensed Electron Configuration | Number of Valence Electrons |
|--------------|---------------|----------------------------------|-----------------------------|
| Sulfur, S    | 16            |                                  |                             |
| Scandium, Sc | 21            |                                  |                             |
| Cesium, Cs   | 55            |                                  |                             |

4. How many valence electrons are in each of the following elements:

Bi (83) \_\_\_\_\_

Ca (20) \_\_\_\_\_

Kr (36) \_\_\_\_\_

As (33) \_\_\_\_\_

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