# Stoichiometry - Mixed Conversions Answer Key

Use the equation below the answer the following questions:

d. How many grams of NH, are made from 31L of N, gas?

e. How many molecules of N, are needed to produce 65g of NH,7

```
\frac{\text{obj. no.}}{1-2} = \frac{1 \cdot \text{not. N_s}}{11 \cdot \text{lights}} = \frac{1 \cdot \text{not. N_s}}{1 \cdot \text{not. N_s}} = \frac{1 \cdot 22 \times 10^{14} \cdot \text{moderators} \cdot N_s}{1 \cdot \text{not. N_s}} = 1 \cdot 22 \times 10^{14} \cdot \text{moderators} \cdot N_s
```

t. How many motes of N, are needed to completely react with 87L of H, gas?

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2. Buttane reacts with oxygen to form water and certion dioxide.

c. How many molecules of HuO are produced with 20.1s of O<sub>0</sub> gas that are completely used sturing the combustion of butane?

 $\frac{10 \times 10^{-1} \, \mathrm{g}^{-1} \, \mathrm{mod} \, R_{\mathrm{p}}}{2 \, \mathrm{s}^{-1} \, \mathrm{s}^{-1} \, \mathrm{mod} \, R_{\mathrm{p}}} = \frac{10 \times 10^{-1} \, \mathrm{mod} \, R_{\mathrm{p}}}{10 \times 10^{-1} \, \mathrm{mod} \, \mathrm{s}^{-1}} = 4.17 \, \mathrm{g} \, 10^{-1} \, \mathrm{mod} \, \mathrm{cube} \, \mathrm{cube} \, R_{\mathrm{p}} \, \mathrm{s}^{-1}$ 

 If 48g of HUD are collected from the burning of bulane, then how many moles of bulane were reacted with oxygen?

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3. POtassium reacts with water to produce potassium hydroxide and hydrogen gas.

c. How many grams of H<sub>i</sub>, will be produced if 7.22 x 10<sup>14</sup> atoms of potassium react completely with water?

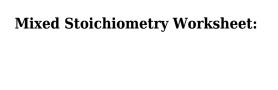
 $\frac{1 \cdot 20 \times 10^{11} \cdot 20 \times 10^{11} \cdot 10^{11}$ 

 How many atoms of K are needed to completely react with 3.22 x 10<sup>th</sup> molecules of water?

\$150 x 10<sup>10</sup> materials (E.D. 2 | 1 marks | 1 marks | 2 marks | 2

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