

4. A solution which contains only one of the following cations: Ca^{2+} , Cu^+ , or Sr^{2+} is tested with various reagents and the following results are obtained:

Reagent	Results
0.2M Na_2SO_4	no precipitate
0.2M NaBr	precipitate
0.2M NaNO_3	no precipitate

Which cation does the solution contain? Cu^+ (1 mark)

5. A solution contains Fe^{3+} , Ca^{2+} , Cu^+ , and Be^{2+} . What compounds (give correct formula for each compound) could be added, and in what order, to separate these ions. Answer this question by filling in the following chart: (8 marks)

First compound - NaCl (Br , I) (Na_2SO_4)

Net-ionic equation for the precipitation reaction: $\text{Cu}^+ + \text{Cl}^- \rightarrow \text{CuCl}(s)$ ($\text{Ca}^{2+} + \text{SO}_4^{2-} \rightarrow \text{CaSO}_4$)

Second compound - Na_2SO_4 (NaI) (Na_2S)

Net-ionic equation for the precipitation reaction: $\text{Ca}^{2+} + \text{SO}_4^{2-} \rightarrow \text{CaSO}_4$ (CuI , Fe_2S_3)

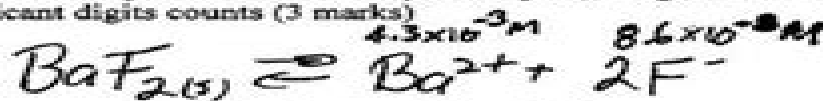
Third compound - Na_2S

Net-ionic equation for the precipitation reaction: $2\text{Fe}^{3+} + 3\text{S}^{2-} \rightarrow \text{Fe}_2\text{S}_3(s)$

Fourth compound - Na_3PO_4 (NaOH) (Na_2CO_3) (Na_2SO_3)

Net-ionic equation for the precipitation reaction: $3\text{Be}^{2+} + 2\text{PO}_4^{3-} \rightarrow \text{Be}_3(\text{PO}_4)_2(s)$

6. A saturated solution of BaF_2 has a $[\text{F}^-]$ of $8.6 \times 10^{-3}\text{M}$ at a certain temperature. Calculate the K_{sp} at this temperature. Show all your steps in a logical manner. Correct use of units and significant digits counts (3 marks)



$$K_{sp} = [\text{Ba}^{2+}][\text{F}^-]^2 = (4.3 \times 10^{-3})(8.6 \times 10^{-3})^2 = 3.18 \times 10^{-7}$$

Answer $K_{sp} = 3.2 \times 10^{-7}$

Mixed Review Chemistry Answer Sheet

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Table of Contents Mixed Review Chemistry Answer Sheet

1. Understanding the eBook Mixed Review Chemistry Answer Sheet
 - The Rise of Digital Reading Mixed Review Chemistry Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Review Chemistry Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Review Chemistry Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Review Chemistry Answer Sheet
 - Personalized Recommendations
 - Mixed Review Chemistry Answer Sheet User Reviews and Ratings

- Mixed Review Chemistry Answer Sheet and Bestseller Lists
- 5. Accessing Mixed Review Chemistry Answer Sheet Free and Paid eBooks
 - Mixed Review Chemistry Answer Sheet Public Domain eBooks
 - Mixed Review Chemistry Answer Sheet eBook Subscription Services
 - Mixed Review Chemistry Answer Sheet Budget-Friendly Options
- 6. Navigating Mixed Review Chemistry Answer Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Review Chemistry Answer Sheet Compatibility with Devices
 - Mixed Review Chemistry Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Review Chemistry Answer Sheet
 - Highlighting and Note-Taking Mixed Review Chemistry Answer Sheet
 - Interactive Elements Mixed Review Chemistry Answer Sheet
- 8. Staying Engaged with Mixed Review Chemistry Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Review Chemistry Answer Sheet
- 9. Balancing eBooks and Physical Books Mixed Review Chemistry Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Review Chemistry Answer Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Review Chemistry Answer Sheet
 - Setting Reading Goals Mixed Review Chemistry Answer Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Review Chemistry Answer Sheet
 - Fact-Checking eBook Content of Mixed Review Chemistry Answer Sheet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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