

MIXED GAS LAWS WORKSHEET

Key

Directions: Examine each question and then write the form of the gas law you plan to use to solve each question. Show which values you have, which values are missing and/or which values need to be calculated. Be careful to use standard units of volume (liters), temperature (Kelvins) & pressure (Atm or mm of Hg).

Note: 1 atm = 760 mm Hg

1. A gas occupies 3.5L at 2.5 mm Hg pressure. What is the volume at 10 mm Hg at the same temperature?

$$P_1 = 2.5 \text{ mmHg} \quad P_2 = 10 \text{ mmHg} \quad P_1 V_1 = P_2 V_2 \quad \frac{(2.5 \text{ mmHg})(3.5 \text{ L})}{10 \text{ mmHg}} = \frac{(10 \text{ mmHg}) V_2}{10 \text{ mmHg}}$$

$$V_1 = 3.5 \text{ L} \quad V_2 = ?$$

$$0.875 \text{ L} = V_2$$

2. A constant volume of oxygen is heated from 100°C to 185°C. The initial pressure is 4.1 atm. What is the final pressure?

$$P_1 = 4.1 \text{ atm} \quad \frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{4.1 \text{ atm}}{373 \text{ K}} = \frac{P_2}{458 \text{ K}}$$

$$T_1 = 100^\circ\text{C} + 273 = 373 \text{ K}$$

$$P_2 = ?$$

$$T_2 = 185^\circ\text{C} + 273 = 458 \text{ K}$$

$$P_2 = \frac{(4.1 \text{ atm})(458 \text{ K})}{373 \text{ K}} = 5.0 \text{ atm}$$

3. A sample of 25L of NH_3 gas at 10°C is heated at constant pressure until it fills a volume of 50L. What is the new temperature in °C?

$$V_1 = 25 \text{ L} \quad \frac{25 \text{ L}}{283 \text{ K}} = \frac{50 \text{ L}}{T_2} \quad T_2 = \frac{(50 \text{ L})(283)}{25 \text{ L}} = 566 \text{ K}$$

$$T_1 = 10^\circ\text{C} + 273 = 283 \text{ K}$$

$$V_2 = 50 \text{ L}$$

$$T_2 = ?$$

$$^\circ\text{C} = \text{K} - 273 = 566 - 273 = 293^\circ\text{C}$$

4. An unknown gas weighs 34g and occupies 6.7L at 2 atm and 245K. What is its molecular weight?

$$\text{Molar Mass} = \frac{g}{\text{mol}} \quad PV = nRT$$

$$(2 \text{ atm})(6.7 \text{ L}) = n (0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})$$

$$n = \frac{(2 \text{ atm})(6.7 \text{ L})}{(0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})} = 0.666 \text{ mol}$$

$$\text{Molar mass} = \frac{34 \text{ g}}{0.666 \text{ mol}} = 51 \text{ g/mol}$$

5. An ideal gas occupies 400ml at 270 mm Hg and 65°C. If the pressure is changed to 1.4 atm and the temperature is increased to 100°C, what is the new volume?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(0.355 \text{ atm})(400 \text{ mL})}{318 \text{ K}} = \frac{(1.4 \text{ atm})(V_2)}{373 \text{ K}}$$

$$V_2 = \frac{(0.355 \text{ atm})(400 \text{ mL})(373 \text{ K})}{(318 \text{ K})(1.4 \text{ atm})} = 112 \text{ mL}$$

$$\text{Unit: } \begin{cases} P_1 = 270 \text{ mmHg} \\ V_1 = 400 \text{ mL} \\ T_1 = 65^\circ\text{C} + 273 = 338 \text{ K} \\ P_2 = 1.4 \text{ atm} \\ V_2 = ? \\ T_2 = 100^\circ\text{C} + 273 = 373 \text{ K} \end{cases}$$

6. What is the volume of 23g of neon gas at 1°C and a pressure of 2 atm?

$$PV = nRT \quad 23 \text{ g Ne} \left(\frac{1 \text{ mol Ne}}{20.18 \text{ g Ne}} \right) = 1.14 \text{ mol Ne}$$

$$V = \frac{nRT}{P} = \frac{(1.14 \text{ mol})(0.0821 \frac{\text{L atm}}{\text{mol K}})(274 \text{ K})}{2 \text{ atm}} = 12.8 \text{ L}$$

Mixed Gas Law Calculations Answers

G Psacharopoulos



Mixed Gas Law Calculations Answers:

This book delves into Mixed Gas Law Calculations Answers. Mixed Gas Law Calculations Answers is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Mixed Gas Law Calculations Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Mixed Gas Law Calculations Answers
- Chapter 2: Essential Elements of Mixed Gas Law Calculations Answers
- Chapter 3: Mixed Gas Law Calculations Answers in Everyday Life
- Chapter 4: Mixed Gas Law Calculations Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Mixed Gas Law Calculations Answers. This chapter will explore what Mixed Gas Law Calculations Answers is, why Mixed Gas Law Calculations Answers is vital, and how to effectively learn about Mixed Gas Law Calculations Answers.
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4. In chapter 3, the author will examine the practical applications of Mixed Gas Law Calculations Answers in daily life. The third chapter will showcase real-world examples of how Mixed Gas Law Calculations Answers can be effectively utilized in everyday scenarios.
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