

24.9 mol

2. What volume would the  $\text{CO}_2$  in Problem 1 occupy at STP?

$$24.9 \text{ mol} \times 22.4 = 557.76 \text{ L}$$

560 L

3. A sample of  $\text{NH}_3$  gas occupies 75.0 liters at STP. How many molecules is this?

$$\frac{75.0 \text{ L}}{22.4} = 3.3482 \text{ mol} \times (6.02 \times 10^{23}) = 20.156 \times 10^{23}$$

4. What is the mass of the sample of  $\text{NH}_3$  in Problem 3?

5. How many atoms are there in  $1.3 \times 10^{22}$  molecules of  $\text{NO}_2$ ?

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